

(FILE 'HOME' ENTERED AT 14:13:31 ON 31 MAY 2003)

FILE 'CAPLUS, WPIDS, COMPENDEX' ENTERED AT 14:13:58 ON 31 MAY 2003

L1 37232 FILE CAPLUS
L2 17098 FILE WPIDS
L3 7915 FILE COMPENDEX
TOTAL FOR ALL FILES
L4 62245 S (FUEL CELL?)
L5 49672 FILE CAPLUS
L6 18949 FILE WPIDS
L7 7478 FILE COMPENDEX
TOTAL FOR ALL FILES
L8 76099 S (NAFION? OR IONOMER? OR ((ION EXCHANGE?)) (10A) (POLYMER OR RE
L9 1915 FILE CAPLUS
L10 448 FILE WPIDS
L11 344 FILE COMPENDEX
TOTAL FOR ALL FILES
L12 2707 S L4 AND L8
L13 7 FILE CAPLUS
L14 1 FILE WPIDS
L15 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L16 8 S L12 AND ((HYDROPHOBIC? OR APOLAR? OR NONPOLAR? OR (NON-POLAR?
L17 86 FILE CAPLUS
L18 35 FILE WPIDS
L19 17 FILE COMPENDEX
TOTAL FOR ALL FILES
L20 138 S L12 AND (INK? OR SCREEN? OR PRINT?)
L21 308 FILE CAPLUS
L22 88 FILE WPIDS
L23 69 FILE COMPENDEX
TOTAL FOR ALL FILES
L24 465 S L12 AND (PRESS? OR SQUEEZ?)
L25 1 FILE CAPLUS
L26 0 FILE WPIDS
L27 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L28 1 S (L16 AND L20 AND L24) — priority
L29 6 FILE CAPLUS
L30 1 FILE WPIDS
L31 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L32 7 S L16 NOT L28
L33 7 DUP REM L32 (0 DUPLICATES REMOVED)
L34 7 FOCUS L33 1-
L35 16 FILE CAPLUS
L36 20 FILE WPIDS
L37 1 FILE COMPENDEX
TOTAL FOR ALL FILES
L38 37 S L20 AND (PRESS? OR SQUEEZ?)
L39 14 FILE CAPLUS
L40 8 FILE WPIDS
L41 1 FILE COMPENDEX
TOTAL FOR ALL FILES
L42 23 S L38 AND (PT OR PLATINUM?)
L43 20 DUP REM L42 (3 DUPLICATES REMOVED)
L44 20 FOCUS L43 1-
L45 14 S L44
L46 6 S L34
L47 14 FILE CAPLUS
L48 6 S L44
L49 1 S L34

L50 6 FILE WPIDS
L51 0 S L44
L52 0 S L34
L53 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L54 20 S L44 NOT L34
L55 20 FOCUS L54 1-
E WILSON M S/AU
L56 23 FILE CAPLUS
L57 26 FILE WPIDS
L58 11 FILE COMPENDEX
TOTAL FOR ALL FILES
L59 60 S E3-E4
L60 11 FILE CAPLUS
L61 12 FILE WPIDS
L62 8 FILE COMPENDEX
TOTAL FOR ALL FILES
L63 31 S L59 AND (FUEL CELL?)
L64 1 FILE CAPLUS
L65 5 FILE WPIDS
L66 1 FILE COMPENDEX
TOTAL FOR ALL FILES
L67 7 S L63 AND PRESS?
L68 7 FOCUS L67 1-
E GOTTESFELD S/AU
L69 197 FILE CAPLUS
L70 21 FILE WPIDS
L71 86 FILE COMPENDEX
TOTAL FOR ALL FILES
L72 304 S E3-E7
L73 97 FILE CAPLUS
L74 12 FILE WPIDS
L75 38 FILE COMPENDEX
TOTAL FOR ALL FILES
L76 147 S L72 AND (FUEL CELL?)
L77 6 FILE CAPLUS
L78 2 FILE WPIDS
L79 2 FILE COMPENDEX
TOTAL FOR ALL FILES
L80 10 S L76 AND PRESS?
L81 8 DUP REM L80 (2 DUPLICATES REMOVED)
L82 6 S L81
L83 1 S L68
L84 5 FILE CAPLUS
L85 2 S L81
L86 5 S L68
L87 2 FILE WPIDS
L88 0 S L81
L89 1 S L68
L90 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L91 7 S L81 NOT L68
L92 7 FOCUS L91 1-
E WILSON MAHLON SCOTT/AU
L93 40 FILE CAPLUS
L94 0 FILE WPIDS
L95 8 FILE COMPENDEX
TOTAL FOR ALL FILES
L96 48 S E1-E3
L97 34 FILE CAPLUS
L98 0 FILE WPIDS
L99 7 FILE COMPENDEX
TOTAL FOR ALL FILES

L100 41 S L96 AND (FUEL CELL?)
L101 4 FILE CAPLUS
L102 0 FILE WPIDS
L103 0 FILE COMPENDEX
TOTAL FOR ALL FILES
L104 4 S L100 AND PRESS?
L105 4 FOCUS L104 1-

FILE 'USPATFULL, USPAT2' ENTERED AT 14:45:10 ON 31 MAY 2003
L106 240 FILE USPATFULL
L107 4 FILE USPAT2
TOTAL FOR ALL FILES
L108 244 S 427/370000/NCL
L109 2 FILE USPATFULL
L110 0 FILE USPAT2
TOTAL FOR ALL FILES
L111 2 S L108 AND (FUEL CELL?)
L112 742 FILE USPATFULL
L113 21 FILE USPAT2
TOTAL FOR ALL FILES
L114 763 S 429/012000/NCL OR 429/042000/NCL
L115 665 FILE USPATFULL
L116 21 FILE USPAT2
TOTAL FOR ALL FILES
L117 686 S L114 AND (FUEL CELL?)
L118 575 FILE USPATFULL
L119 19 FILE USPAT2
TOTAL FOR ALL FILES
L120 594 S L117 AND (PRESS? OR SQUEEZ?)
L121 12 FILE USPATFULL
L122 1 FILE USPAT2
TOTAL FOR ALL FILES
L123 13 S L120 AND ((HYDROPHOBIC? OR NONPOLAR? OR APOLAR? OR (NON-POLAR
L124 13 FOCUS L123 1-

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L34 ANSWER 1 OF 7 CAPLUS COPYRIGHT 2003 ACS
 AN 1996:388536 CAPLUS
 DN 125:38070
 TI Manufacture of electrodes for solid polymer electrolyte **fuel cells**
 IN Tada, Tomoyuki
 PA Tanaka Precious Metal Ind, Japan; Watanabe Masahiro; Sutonharuto Asosheetsu Inc
 SO Jpn. Kokai Tokkyo Koho, 5 pp.
 CODEN: JKXXAF
 DT Patent
 LA Japanese
 IC ICM H01M004-88
 ICS B01J037-00; B01J037-02; H01M004-86; H01M008-02; H01M008-10
 ICA B01J023-42
 CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
 FAN.CNT 4

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 08115726	A2	19960507	JP 1994-277108	19941017
	US 5843519	A	19981201	US 1995-543632	19951016
PRAI	JP 1994-277108		19941017		
	JP 1994-332291		19941017		
	JP 1994-289288		19941028		
	JP 1994-289289		19941028		

AB The electrodes are prep'd. by spray drying a dispersion of ground catalyst particles in an org. **solvent**, contg. **ion exchanger resin** and optionally a **hydrophobic resin**, to obtain **resin** coated catalyst granules and applying the granules on a substrate to form a catalyst layer. Preferably, the ground catalyst particles have diam. 0.1-10 .mu.m, the granules have diam. 1-50 .mu.m, the dispersion contains 0.5-15% solids, the spraying is carried out at 90-160.degree. and 0.8-1.5 kg/cm² spraying pressure, and the solvent has b. ltoreq. 160.degree..

ST solid polymer electrolyte **fuel cell** electrode; **fuel cell** electrode catalyst resin coating; electrode catalyst ion exchanger coating; hydrophobic resin coating electrode catalyst

IT Polyoxyalkylenes, uses
 RL: NUU (Other use, unclassified); USES (Uses)
 (fluorine- and sulfo-contg., **ionomers**, manuf. of **Nafion** coated catalyst granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

IT Electrodes
 (fuel-cell, manuf. of **Nafion** coated catalyst granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

IT Fluoropolymers
 RL: NUU (Other use, unclassified); USES (Uses)
 (polyoxyalkylene-, sulfo-contg., **ionomers**, manuf. of **Nafion** coated catalyst granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

IT **Ionomers**
 RL: NUU (Other use, unclassified); USES (Uses)
 (polyoxyalkylenes, fluorine- and sulfo-contg., manuf. of **Nafion** coated catalyst granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

IT Drying
 (spray, spray drying in manuf. of **Nafion** coated catalyst)

granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

IT 7440-06-4, Platinum, uses 7440-44-0, Carbon, uses
RL: CAT (Catalyst use); PEP (Physical, engineering or chemical process);
PROC (Process); USES (Uses)
(manuf. of **Nafion** coated catalyst granules contg. platinum loaded on carbon support for solid polymer electrolyte **fuel cell** electrodes)

RN 7440-06-4

RN 7440-44-0

L34 ANSWER 2 OF 7 CAPLUS COPYRIGHT 2003 ACS

AN 1994:659704 CAPLUS

DN 121:259704

TI Manufacture of solid polymer electrolyte **fuel cells**

IN Seki, Tsutomu

PA Tokyo Gas Co Ltd, Japan

SO Jpn. Kokai Tokkyo Koho, 4 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

IC ICM H01M008-02

ICS H01M004-86; H01M004-88; H01M008-10

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)

FAN. CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 06203849	A2	19940722	JP 1992-358059	19921225
PRAI	JP 1992-358059		19921225		

AB The **fuel cells** are prep'd. by mixing carbon black loaded Pt catalyst and a **ion exchanger resin** used as solid **polymer** electrolyte in a **solvent** to form a suspension, depositing the suspension on **hydrophobically** treated substrates to form electrode sheets, holding an **ion exchanger polymer** electrolyte membrane between an electrode sheet pair, and hot pressing.

ST solid polymer electrolyte **fuel cell**; polymer electrolyte **fuel cell** manuf

IT **Fuel cells**

(manuf. of solid polymer electrolyte **fuel cells**)

IT Carbon black, uses

RL: CAT (Catalyst use); DEV (Device component use); USES (Uses)
(manuf. of solid polymer electrolyte **fuel cells**)

IT 7440-06-4, Platinum, uses

RL: CAT (Catalyst use); DEV (Device component use); USES (Uses)
(manuf. of solid polymer electrolyte **fuel cells**)

IT 66796-30-3, **Nafion** 117

RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)

(manuf. of solid polymer electrolyte **fuel cells**)

RN 7440-06-4

RN 66796-30-3

L34 ANSWER 3 OF 7 CAPLUS COPYRIGHT 2003 ACS

AN 1994:659705 CAPLUS

DN 121:259705

TI Manufacture of solid polymer electrolyte **fuel cells**

IN Seki, Tsutomu

PA Tokyo Gas Co Ltd, Japan

SO Jpn. Kokai Tokkyo Koho, 4 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

have

IC ICM H01M008-02
ICS H01M004-86; H01M004-88; H01M008-10
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 06203848	A2	19940722	JP 1992-358058	19921225
PRAI	JP 1992-358058		19921225		
AB	The fuel cells are prep'd. by mixing carbon black loaded Pt catalyst and a ion exchanger resin used as solid polymer electrolyte in a solvent to form a slurry, applying the slurry to a hydrophobically treated electrode substrate, removing the solvent by evapn. to form an electrode sheet, and hot pressing an ion exchanger membrane between a pair of the electrode sheets.				
ST	solid polymer electrolyte fuel cell; polymer electrolyte fuel cell manuf				
IT	Fuel cells (manuf. of solid polymer electrolyte fuel cells)				
IT	Carbon black, uses RL: CAT (Catalyst use); DEV (Device component use); USES (Uses) (manuf. of solid polymer electrolyte fuel cells)				
IT	7440-06-4, Platinum, uses RL: CAT (Catalyst use); DEV (Device component use); USES (Uses) (manuf. of solid polymer electrolyte fuel cells)				
IT	66796-30-3, Nafion 117 RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses) (manuf. of solid polymer electrolyte fuel cells)				
RN	7440-06-4				
RN	66796-30-3				

L34 ANSWER 4 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 1999:392852 CAPLUS
DN 131:33839

TI **Fuel cell** electrodes and their manufacture
IN Yamada, Hiroshi
PA Tokyo Gas Co., Ltd., Japan
SO Jpn. Kokai Tokkyo Koho, 6 pp.
CODEN: JKXXAF

DT Patent
LA Japanese
IC ICM H01M004-86
ICS H01M004-88; H01M008-10

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 11167925	A2	19990622	JP 1997-365849	19971222
PRAI	JP 1997-284428		19971001		
AB	The electrodes have a catalyst layer, formed on a porous gas diffusion layer, and are prep'd. by evapg. a solvent from a suspension contg. catalyst particles, an electrolyte, and a hydrophobic agent under controlled temp. and pressure to maintain an evapn. rate of .apprx.8 cm ³ /min. The catalyst is preferably Pt, Pd, and/or their alloy loaded on C particles; the electrolyte is a perfluorocarbon sulfonic acid; the hydrophobic agent is polytetrafluoroethylene; the solvent is water and/or alc.; and the fuel cells are polymer electrolyte fuel cells.				
ST	fuel cell electrode catalyst layer manuf; evapn control fuel cell electrode manuf				
IT	Evaporation Fuel cell electrodes				

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(controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 IT Carbon black, uses
 RL: CAT (Catalyst use); USES (Uses)
 (controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 IT Fluoropolymers, uses
 RL: DEV (Device component use); USES (Uses)
 (controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 IT 7440-06-4, Platinum, uses
 RL: CAT (Catalyst use); USES (Uses)
 (controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 IT 9002-84-0, Polytetrafluoroethylene 66796-30-3, **Nafion** 117
 RL: DEV (Device component use); USES (Uses)
 (controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 IT 64-17-5, Ethanol, processes 7732-18-5, Water, processes
 RL: REM (Removal or disposal); PROC (Process)
 (controlled evapn. of solvents in manuf. of catalyst layers for polymer
 electrolyte **fuel cell** electrodes)
 RN 7440-06-4
 RN 9002-84-0
 RN 66796-30-3
 RN 64-17-5
 RN 7732-18-5

L34 ANSWER 5 OF 7 CAPLUS COPYRIGHT 2003 ACS
 AN 1999:631497 CAPLUS
 DN 131:259912
 TI Membrane electrode assembly for polymer electrolyte membrane **fuel**
cell and method for its manufacture
 IN Zuber, Ralf; Fehl, Knut; Starz, Karl-anton; Stenke, Udo
 PA Degussa-Huls A.-G., Germany
 SO Eur. Pat. Appl., 13 pp.
 CODEN: EPXXDW



DT Patent
 LA German
 IC ICM H01M008-10
 ICS H01M004-92
 CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
 Section cross-reference(s): 37, 67

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 945910	A2	19990929	EP 1999-104630	19990309
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	DE 19812592	A1	19991007	DE 1998-19812592	19980323
	US 6309772	B1	20011030	US 1999-274018	19990322
	JP 11329452	A2	19991130	JP 1999-77861	19990323
	BR 9900605	A	20000606	BR 1999-605	19990323

PRAI DE 1998-19812592 A 19980323

AB The membrane electrode assembly of the **fuel cell**
 comprises a polymer electrolyte membrane with porous reaction layers
 contg. catalysts and **ionomers** on both sides of the membrane.
 The reaction layer has an inhomogeneous microstructure formed from an
ionomer-impregnated and embedded catalyst portion and an
ionomer-free catalyst portion in wt. ratio (1-20):1, esp.
 (3-10):1. The catalyst can be carbon-supported Pt-group metal or alloy
 particles. The reaction layer has pore vol. 0.7-1.3, esp. 0.8-1.2 mL/g,
 for pores with diam. 0.03-1 .mu.m, and thickness 5-100, esp. 10-100 .mu.m.

The ionomer can be a proton-conducting tetrafluoroethylene-fluorovinylether copolymer contg. acid groups, e.g., **Nafion**.

ST membrane electrode assembly PEM **fuel cell**; polymer electrolyte membrane **fuel cell**

IT Carbon black, uses
RL: CAT (Catalyst use); USES (Uses)
(catalyst supports; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Platinum-group metals
RL: CAT (Catalyst use); USES (Uses)
(catalysts; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Glycols, uses
RL: NUU (Other use, unclassified); USES (Uses)
(ethers, solvents; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Polyoxyalkylenes, uses
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(fluorine- and sulfo-contg., **ionomers**, proton-conducting; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Polyoxyalkylenes, uses
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(fluorine-contg., sulfo-contg., **ionomers**, proton-conducting; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Ethers, uses
RL: NUU (Other use, unclassified); USES (Uses)
(glycol, solvents; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT **Fuel cell** electrolytes
(polymer membranes; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT **Fuel cells**
(polymer-electrolyte-membrane; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Fluoropolymers, uses
Fluoropolymers, uses
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(polyoxyalkylene-, sulfo-contg., **ionomers**, proton-conducting; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT **Ionomers**
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(polyoxyalkylenes, fluorine- and sulfo-contg., proton-conducting; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Fluoropolymers, uses
Ionomers
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(proton-conducting; membrane electrode assembly for polymer electrolyte membrane **fuel cells**)

IT Alcohols, uses
Glycols, uses
Hydrocarbons, uses
Paraffin oils
RL: NUU (Other use, unclassified); USES (Uses)
(solvents; membrane electrode assembly for polymer electrolyte membrane

IT fuel cells)
IT Solvents
 (weakly polar; nonpolar; membrane electrode assembly for
 polymer electrolyte membrane fuel cells)
IT 7440-05-3, Palladium, uses 7440-06-4, Platinum, uses 7440-16-6,
 Rhodium, uses
 RL: CAT (Catalyst use); USES (Uses)
 (catalysts; membrane electrode assembly for polymer electrolyte
 membrane fuel cells)
IT 77950-55-1, Nafion 115
 RL: DEV (Device component use); TEM (Technical or engineered material
 use); USES (Uses)
 (membranes; membrane electrode assembly for polymer electrolyte
 membrane fuel cells)
IT 7439-89-6, Iron, uses 7439-98-7, Molybdenum, uses 7440-02-0, Nickel,
 uses 7440-18-8, Ruthenium, uses 7440-33-7, Tungsten, uses 7440-47-3,
 Chromium, uses 7440-48-4, Cobalt, uses 7440-50-8, Copper, uses
 7440-62-2, Vanadium, uses
 RL: CAT (Catalyst use); USES (Uses)
 (platinum group metals alloyed with, catalysts; membrane electrode
 assembly for polymer electrolyte membrane fuel cells
)
IT 116-14-3D, Tetrafluoroethylene, fluorovinylether copolymers,
 functionalized 57578-63-9D, Perfluorovinylether-tetrafluoroethylene
 copolymer, functionalized
 RL: DEV (Device component use); TEM (Technical or engineered material
 use); USES (Uses)
 (proton-conducting; membrane electrode assembly for polymer electrolyte
 membrane fuel cells)
IT 56-81-5, 1,2,3-Propanetriol, uses 57-55-6, 1,2-Propanediol, uses
 107-41-5, Hexylene glycol 110-38-3, Decanoic acid, ethyl ester
 111-82-0, Dodecanoic acid, methyl ester 463-79-6D, Carbonic acid, alkyl
 esters, uses 25265-71-8, Dipropylene glycol
 RL: NUU (Other use, unclassified); USES (Uses)
 (solvents; membrane electrode assembly for polymer electrolyte membrane
 fuel cells)
RN 7440-05-3
RN 7440-06-4
RN 7440-16-6
RN 77950-55-1
RN 7439-89-6
RN 7439-98-7
RN 7440-02-0
RN 7440-18-8
RN 7440-33-7
RN 7440-47-3
RN 7440-48-4
RN 7440-50-8
RN 7440-62-2
RN 116-14-3D
RN 57578-63-9D
RN 56-81-5
RN 57-55-6
RN 107-41-5
RN 110-38-3
RN 111-82-0
RN 463-79-6D
RN 25265-71-8

L34 ANSWER 6 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 2001:78453 CAPLUS
DN 134:148631
TI Fluid composition for producing and repairing ion-exchange membranes

IN Blach Vizoso, Ricardo; Timofeev, Sergei; Bobrova, Lyubov; Fateev, Vladimir
PA David Systems Technology, S.L., Spain
SO PCT Int. Appl., 29 pp.
CODEN: PIXXD2
DT Patent
LA Spanish
IC ICM C08L027-18
 ICS C08L027-12
CC 38-3 (Plastics Fabrication and Uses)
 Section cross-reference(s): 48, 52, 67, 72
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2001007517	A1	20010201	WO 1999-ES278	19990827
	·W: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM ·RW: GH, GM, KE, LS, MW, SD, SL, SZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	ES 2154231	A1	20010316	ES 1999-1653	19990722
	ES 2154231	B1	20011201		
	AU 9957473	A1	20010213	AU 1999-57473	19990827
	EP 1209197	A1	20020529	EP 1999-944645	19990827
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO, MK, CY, AL				
	BR 9917416	A	20020702	BR 1999-17416	19990827
	JP 2003505569	T2	20030212	JP 2001-512793	19990827
PRAI	ES 1999-1653	A	19990722		
	WO 1999-ES278	W	19990827		

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AB The fluid compn. contains 1-35 wt.% of a perfluorinated ion-exchange copolymer with functional groups SO₃M (M = H, alkali metal) having equiv. wt. >900 and 65-99% of a polar org. **solvent** or its mixt. with a **nonpolar solvent**. The perfluorinated ion-exchange copolymer has crystallinity 2-10% and the ratio of its d. to that of the fluoropolymer without sulfo groups is 0.9-0.97. Such compns. are used in the prodn. and repair of ion-exchange membranes which are used in electrolysis for manuf. of chlorine and caustic or in water electrolysis in fuel/gas sepn. cells. A 2% soln. of SPL 2 [84:16 copolymer of tetrafluoroethylene with CF₂:CFOCF₂CF(CF₃)OCF₂CF₂SO₃H] in iso-PrOH was used in the manuf. of a **fuel cell** electrode.

ST **ion exchange resin repair soln;**
ionomer perfluoro sulfonated ion exchanger; fuel cell membrane cation exchanger

IT Cation exchange membranes
(fluid compn. for producing and repairing ion-exchange membranes)

IT Fluoropolymers, uses
RL: DEV (Device component use); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)
(fluid compn. for producing and repairing ion-exchange membranes)

IT Dealkylation catalysts
(fluid compn. for producing and repairing ion-exchange membranes for use as)

IT Electrolysis
Fuel cells

IT Hygrometers
(fluid compn. for producing and repairing ion-exchange membranes for use in)

IT Membranes, nonbiological
(semipermeable; fluid compn. for producing and repairing ion-exchange

membranes for use as)
IT 128-37-0, 2,6-Di-tert-butyl-4-methylphenol, reactions
RL: RCT (Reactant); RACT (Reactant or reagent)
(dealkylation catalysts for)
IT 2409-55-4P, 2-tert-Butyl-4-methylphenol
RL: IMF (Industrial manufacture); PREP (Preparation)
(dealkylation catalysts for manuf. of)
IT 31176-88-2, SPL 1
RL: CAT (Catalyst use); PEP (Physical, engineering or chemical process);
PROC (Process); USES (Uses)
(fluid compn. for producing and repairing ion-exchange membranes)
IT 31175-20-9, SPL 2 278616-83-4, SPL 4 324011-45-2, SPL 6
RL: DEV (Device component use); PEP (Physical, engineering or chemical
process); PROC (Process); USES (Uses)
(fluid compn. for producing and repairing ion-exchange membranes)
IT 64-17-5, Ethanol, uses 67-56-1, Methanol, uses 67-63-0, Isopropyl
alcohol, uses 67-64-1, Acetone, uses 67-68-5, DMSO, uses 68-12-2,
DMF, uses 71-23-8, 1-Propanol, uses 71-43-2, Benzene, uses 78-83-1,
Isobutyl alcohol, uses 78-93-3, Methyl ethyl ketone, uses 108-88-3,
Toluene, uses 108-94-1, Cyclohexanone, uses 109-66-0, Pentane, uses
110-54-3, Hexane, uses 127-19-5, N,N-Dimethylacetamide 142-82-5,
Heptane, uses 594-41-2, 2-Bromo-1,1,1-trichloroethane 761-65-9,
N,N-Dibutylformamide
RL: NUU (Other use, unclassified); USES (Uses)
(fluid compn. for producing and repairing ion-exchange membranes)
IT 269080-43-5, SPL 7 324011-44-1, SPL 3
RL: POF (Polymer in formulation); USES (Uses)
(fluid compn. for producing and repairing ion-exchange membranes)
IT 7664-39-3, Hydrogen fluoride, processes
RL: REM (Removal or disposal); PROC (Process)
(membranes for sepn. from halocarbons of)
IT 354-21-2, 1,2,2-Trichloro-1,1-difluoroethane 354-23-4,
1,2-Dichloro-1,1,2-trifluoroethane 1649-08-7, 1,2-Dichloro-1,1-
difluoroethane
RL: NUU (Other use, unclassified); USES (Uses)
(membranes for sepn. of HF from)
IT 76-13-1P, 1,1,2-Trichloro-1,2,2-trifluoroethane
RL: NUU (Other use, unclassified); PUR (Purification or recovery); PREP
(Preparation); USES (Uses)
(membranes for sepn. of HF from)

RE.CNT 3 THERE ARE 3 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Asahi Glass Company Ltd; EP 0025644 A 1981 CAPLUS
- (2) Michael, J; US 4386987 A 1983 CAPLUS
- (3) Walther, G; US 4453991 A 1984 CAPLUS

RN 128-37-0
RN 2409-55-4P
RN 31176-88-2
RN 31175-20-9
RN 278616-83-4
RN 324011-45-2
RN 64-17-5
RN 67-56-1
RN 67-63-0
RN 67-64-1
RN 67-68-5
RN 68-12-2
RN 71-23-8
RN 71-43-2
RN 78-83-1
RN 78-93-3
RN 108-88-3
RN 108-94-1

RN 109-66-0
 RN 110-54-3
 RN 127-19-5
 RN 142-82-5
 RN 594-41-2
 RN 761-65-9
 RN 269080-43-5
 RN 324011-44-1
 RN 7664-39-3
 RN 354-21-2
 RN 354-23-4
 RN 1649-08-7
 RN 76-13-1P

L34 ANSWER 7 OF 7 WPIDS (C) 2003 THOMSON DERWENT
 AN 2001-102762 [11] WPIDS
 DNN N2001-076259 DNC C2001-030126
 TI Electrode-membrane combination for use in fuel cell of transportation vehicle, comprises electrode-membrane interfacial regions comprising zone(s) of specific thickness, which contain catalytically active metals.
 DC A35 A85 L03 M13 X16 X21
 IN ARPS, J H; CAVALCA, C A; MURTHY, M; CAVALCA, C
 PA (GORE) GORE ENTERPRISE HOLDINGS INC; (ARPS-I) ARPS J H; (CAVA-I) CAVALCA C A; (MURT-I) MURTHY M; (GORE) GORE ENTERPRISE HOLDINGS
 CYC 80
 PI WO 2000079630 A2 20001228 (200111)* EN 125p H01M008-10
 RW: AT BE CH CY DE DK ES FI FR GB GR IE IT LU MC NL PT SE
 W: AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB GE
 GH GM HR HU ID IL IS JP KE KG KP KR KZ LC LK LR LS LT LU LV MD MG
 MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR TT UA UG
 UZ VN YU ZW
 AU 2000057440 A 20010109 (200122) H01M008-10
 US 6300000 B1 20011009 (200162) H01M004-86
 US 2001033960 A1 20011025 (200170) H01M004-94
 EP 1201002 A2 20020502 (200236) EN H01M008-10
 R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
 RO SE SI
 CN 1370334 A 20020918 (200303) H01M008-10
 JP 2003502829 W 20030121 (200308) 145p H01M004-86
 ADT WO 2000079630 A2 WO 2000-US16645 20000616; AU 2000057440 A AU 2000-57440 20000616; US 6300000 B1 US 1999-335718 19990618; US 2001033960 A1 Cont of US 1999-335718 19990618, US 2001-827894 20010409; EP 1201002 A2 EP 2000-942876 20000616, WO 2000-US16645 20000616; CN 1370334 A CN 2000-811793 20000616; JP 2003502829 W WO 2000-US16645 20000616, JP 2001-505094 20000616

FDT AU 2000057440 A Based on WO 200079630; EP 1201002 A2 Based on WO 200079630; JP 2003502829 W Based on WO 200079630

PRAI US 1999-335718 19990618; US 2001-827894 20010409

IC ICM H01M004-86; H01M004-94; H01M008-10
 ICS C23C014-14; H01M004-88; H01M004-90; H01M004-92; H01M004-96;
 H01M008-00; H01M008-02

AB WO 200079630 A UPAB: 20010224
 NOVELTY - An electrode-membrane combination comprises reactant diffusive, electronically conductive electrodes (1, 3) comprising catalytically active metal(s) (A) and ionically conductive polymer(s). Ionically conductive membrane(s) (2) contact the electrode to form electrode-membrane interfacial regions (4,5) comprising zone(s) of thickness 3-5000 Angstrom, which contain different catalytically active metals (B).
 DETAILED DESCRIPTION - An INDEPENDENT CLAIM is also included for membrane-electrode assembly comprising the electrode-membrane combination and at least one additional, different electrode contacting the membrane.

USE - For membrane-electrode assembly of **fuel cell** used in transportation vehicle (all claimed) such as car, bus, trucks; stationary power applications; portable power applications such as portable television, fans and other consumer products.

ADVANTAGE - The electrode-membrane combination provides improved power output and resistance to poisoning, during **fuel cell** operation. The ionically conductive membrane provides strength, high ionic conductance and good interfacial contact with the electrode. The membrane damage caused by hot pressing of electrode with membrane is prevented. Inexpensive **fuel cell** of high power density, low cost, high efficiency and durability, is enabled.

DESCRIPTION OF DRAWING(S) - The figure shows the cross-section of membrane-electrode assembly.

Electronically conductive electrodes 1,3

Polymeric membrane 2

Interfacial regions 4,5

Dwg.1/62

FS CPI EPI

FA AB; GI

MC CPI: A12-E06B; A12-T04C; A12-W11K; L03-E04; L03-H05; M13-H05

EPI: X16-C01C; X16-E06A; X16-F02; X21-A01F; X21-B01A

=>

L55 ANSWER 1 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1999:422482 CAPLUS
DN 131:150608
TI Preparation and performance evaluation of membrane electrode assemblies for polymer electrolyte **fuel cell**
AU Peck, D.-H.; Chun, Y.-G.; Kim, C.-S.; Jung, D.-H.; Shin, D.-R.
CS Korea Institute of Energy Research, Taejon, 305-343, S. Korea
SO Journal of New Materials for Electrochemical Systems (1999), 2(2), 121-124
CODEN: JMESFQ; ISSN: 1480-2422
PB Journal of New Materials for Electrochemical Systems
DT Journal
LA English
AB In order to evaluate performance characteristics of the membrane electrode assembly (MEA) for polymer electrolyte **fuel cell** (PEFC), three com. available **Nafion** membranes (112, 115, and 117), and **Pt/C**, **PtNi/C** and **PtNiCo/C** electrocatalysts were used in the fabrication of the MEAs by using transfer **printing** technique. The electrocatalyst layer of the MEA has been made by using a slurry of carbon-supported catalyst, **Nafion-ionomers** and glycerin. The effects of the thickness of **Nafion** membranes, electrocatalysts and the operating conditions (i.e. temp., reactant gas **pressure**, and compn.) on the performance of the MEA were investigated in the PEFC single cell. The performance of the MEA made from **Nafion** 112, 115, and 117 membrane for O₂/H₂ cell were 1.16, 0.82, and 0.54 A/cm² at 0.6V. An increase of reactants **pressures** up to 3 atm has stronger effect on the performance of Air/H₂ cell than on that of O₂/H₂ cell. The performances of the MEA made from **PtNi/C** alloy catalyst for O₂/H₂ and Air/H₂ cells at 80.degree. were 0.86 and 0.48 A/cm² at 0.6V, resp., and in the case of **PtNiCo/C** alloy catalyst, the MEAs for O₂/H₂ and Air/H₂ cells showed similar performance.

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 2 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1998:216367 CAPLUS
DN 129:46606
TI A rapid half-cell technique for the **pre-screening** of polymer **fuel cell** catalysts
AU TamizhMani, G.; Dodelet, J. P.; Guay, D.; Dignard-Bailey, L.
CS Natural Resources Canada, CANMET-Energy Diversification Research Laboratory, Varennes, QC, J3X 1S6, Can.
SO Journal of Electroanalytical Chemistry (1998), 444(1), 121-125
CODEN: JECHE; ISSN: 0368-1874
PB Elsevier Science S.A.
DT Journal
LA English
AB Four platinum-based catalysts with different catalytic activity for the oxygen redn. reaction were prep'd. and tested in polymer **fuel cells** (PFCs) and in half-cells with H₂SO₄ and HF electrolytes. The activity results of PFCs at 0.9 V vs. RHE (reversible hydrogen electrode) can be mimicked in parallel by the results obtained in HF electrolyte but not by the results obtained in H₂SO₄ electrolyte. This paper concludes that the **pre-screening** of a huge no. of **Pt**-based catalysts for the selection of potential catalysts for the PFCs can be carried out by a rapid half-cell technique with a nonadsorbing electrolyte such as HF.

RE.CNT 8 THERE ARE 8 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 3 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1993:84416 CAPLUS

DN 118:84416
TI Membrane catalyst layer for fuel cells
IN Wilson, Mahlon Scott
PA United States Dept. of Energy, USA
SO PCT Int. Appl., 29 pp.
CODEN: PIXXD2
DT Patent
LA English
FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9215121	A1	19920903	WO 1992-US1058	19920218
	W: CA, JP				
	RW: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LU, MC, NL, SE				
	<u>US 656329</u>	A0	19930501	US 1991-656329	19910219
	<u>US 5234777</u>	A	19930810	US 1991-736876	19910729
	JP 05507583	T2	19931028	JP 1992-507039	19920218
	EP 600888	A1	19940615	EP 1992-907218	19920218
	EP 600888	B1	19970827		
	R: DE, FR, GB				
PRAI	US 1991-656329		19910219		✓
	US 1991-736876		19910729		
	WO 1992-US1058		19920218		
AB	<p>Fuel cells incorporate a $\sim 10\text{-}\mu\text{m}$ catalyst layer between a solid polymer electrolyte membrane and a porous electrode backing. The catalyst layer has C-supported Pt catalyst loading >0.1 and $\sim 0.35\text{ mg Pt/cm}^2$. The layer is formed as an ink that is spread and cured on a film-release blank. The cured film is transferred to the membrane and hot pressed into the surface to form a catalyst layer having a controlled thickness and catalyst distribution. Alternately, the catalyst layer is formed by applying a Na^+ form of a perfluorosulfonate ionomer directly to the membrane, drying the film at $\sim 150\text{ degree}$., and then converting the film back to the protonated form of the ionomer. The layer has an adequate gas permeability so that the cell performance is not affected and has an effective d. and particle distribution to optimize proton access to the catalyst and electronic continuity for electron flow from the half-cell reaction occurring at the catalyst.</p>				

L55 ANSWER 4 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1992:87618 CAPLUS
DN 116:87618
TI Thin-film catalyst layers for polymer electrolyte fuel cell electrodes *David*
AU Wilson, M. S.; Gottesfeld, S.
CS Electron. Res. Group, Los Alamos Natl. Lab., Los Alamos, NM, 87545, USA
SO Journal of Applied Electrochemistry (1992), 22(1), 1-7
CODEN: JAELBJ; ISSN: 0021-891X
DT Journal
LA English
AB New structures for the Pt/C catalyst layer of polymer electrolyte fuel cell electrodes have been developed that substantially increase the utilization efficiency of the catalyst. Fabricating the catalyst layers and gas diffusion backings sep. makes it possible to formulate each structure with the properties that are most suitable for its function. In the case of the catalyst layer, the optimal properties are hydrophilicity, thinness, uniformity, and the proper ratio of ionomer and supported catalyst. The catalyst layers are cast from soln. as thin films that utilize the ionomer itself as a binder. The thin films are hot pressed directly onto the ionomer membranes, and the hydrophobic gas diffusion backings are inserted when the cells are assembled. The performances of fuel cells based on the thin film catalyst layers are comparable with

those of gas diffusion electrode designs that utilize several times as much Pt, thus the specific activities of the Pt catalyst in the new structures are significantly higher.

L55 ANSWER 5 OF 20 CAPLUS COPYRIGHT 2003 ACS

AN 1963:473643 CAPLUS

DN 59:73643

OREF 59:13602h,13603a

TI Fuel cell

IN Grubb, Willard T., Jr.

PA Electric Co.

SO 5 pp.

DT Patent

LA Unavailable

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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PI US 2913511 19591117 US 19550629

AB A gaseous fuel cell contg. an ion-exchange resin membrane as the electrolyte and operable at room temp. and atm. pressure is described. The electrodes in the form of sheets or screens are prep'd. from conductive materials which can absorb the fuel employed and act as a catalyst. For example, Amberplex C-1 cation-permeable membrane, 0.025 in. thick, is placed between two 0.0005-in. thick Pt. sheets. The area of contact between the electrodes and the membrane is 25 sq. cm. When H is supplied to the anode and O to the cathode, the cell has an open-circuit voltage of 0.95 v. and delivers a current of 10 ma. at 0.65 v.

L55 ANSWER 6 OF 20 CAPLUS COPYRIGHT 2003 ACS

AN 1999:511087 CAPLUS

DN 131:146953

TI A method of forming a membrane electrode assembly for a direct-feed fuel cell

IN Kindler, Andrew; Dawson, Stephen F.

PA California Institute of Technology, USA

SO PCT Int. Appl., 22 pp.

CODEN: PIXXD2

DT Patent

LA English

FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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PI WO 9939840 A1 19990812 WO 1999-US2835 19990209

W: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM
RW: GH, GM, KE, LS, MW, SD, SZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG

AU 9925961 A1 19990823 AU 1999-25961 19990209

EP 1054739 A1 20001129 EP 1999-905918 19990209

R: DE, FR, GB, NL

PRAI US 1998-21694 A 19980210

WO 1999-US2835 W 19990209

AB A catalyst-coated electrode for a fuel cell is prep'd. by mixing a catalyst (e.g., Pt or Pt-Ru) with a water repellent material (e.g., PTFE) to form a catalyst ink that is applied to an electrode backing material (e.g., porous carbon fiber sheet). The coated electrode is sintered under N₂, cooled to 25.degree.C, then coated with a liq. ionomer (e.g., Nafion--a perfluorovinylether sulfonic acid-tetrafluoroethylene copolymer) forming

an anode or cathode. A solid electrolyte membrane, e.g., a perfluorinated proton exchange membrane, is pretreated (with isopropanol) to soften and swell the membrane prior to hot **press** bonding between the anode and cathode to form a membrane electrode assembly. Swelling the membrane before bonding results in shrinkage at the interface during use, reducing delamination. The electrode assembly can be used in a direct-feed methanol **fuel cell**.

RE.CNT 2 THERE ARE 2 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 7 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1999:511088 CAPLUS
DN 131:146954
TI Direct deposit of catalyst on the membrane of direct-feed **fuel cells**
IN Chun, William; Narayanan, Sekharipuram R.; Jeffries-Nakamura, Barbara; Valdez, Thomas I.; Linke, Juergen
PA California Institute of Technology, USA
SO PCT Int. Appl., 36 pp.
CODEN: PIXXD2
DT Patent
LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9939841	A1	19990812	WO 1999-US2836	19990209
	W: AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM, RW: GH, GM, KE, LS, MW, SD, SZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	US 6221523	B1	20010424	US 1998-21692	19980210
	AU 9925962	A1	19990823	AU 1999-25962	19990209
	EP 1054740	A1	20001129	EP 1999-905919	19990209
	R: DE, FR, GB, NL				
	US 2001052389	A1	20011220	US 2001-933684	20010820
PRAI	US 1998-21692	A	19980210		
	WO 1999-US2836	W	19990209		
	US 1999-428123	A3	19991026		
AB	A solid membrane electrolyte and membrane electrode assembly are prep'd. for a direct liq.-feed fuel cell , resulting in improved catalyst use and an improved catalyst/membrane interface. The catalyst layer is applied directly onto the solid electrolyte membrane as an ink . The catalyst ink is prep'd. by mixing a catalyst (e.g., Pt or Pt-Ru) with a water repellent material (e.g., PTFE) and an ionomer (Nafion) soln. The solid electrolyte membrane, e.g., a perfluorinated proton exchange membrane, is pretreated (with isopropanol) to soften and swell the membrane prior to coating. The pretreated membrane is held in a frame and the catalyst ink is poured or sprayed directly onto the membrane surface and then spread with a glass rod. The coated membrane is dried by slow evapn. Support substrates (e.g., carbon paper) are placed on each side of the catalyst-coated membrane, then the layers are bonded by hot pressing to form the membrane electrode assembly. The membrane electrode assembly can be used in a direct-feed methanol fuel cell .				

RE.CNT 3 THERE ARE 3 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 8 OF 20 CAPLUS COPYRIGHT 2003 ACS

AN 2002:579507 CAPLUS
DN 137:386984
TI In situ voltammetric characterization of PEM fuel cell catalyst layers
AU Kumpulainen, Heikki; Peltonen, Terttu; Koponen, Ulla; Bergelin, Mikael; Valkainen, Matti; Wasberg, Mikael
CS VTT Processes, FIN-02044, Finland
SO VTT Tiedotteita (2002), 2137, 1-28, 1/1-1/4
CODEN: VTIEEE; ISSN: 1235-0605
DT Report
LA English
AB In our work we have studied inhouse made membrane electrode assembly (MEA) (geometric area 5 cm²) based on com. (E-TEK, Inc.) carbon supported Pt and Pt/Ru catalysts. A Nafion 115 membrane was coated with ink contg. Pt/C catalyst and Nafion soln. and dried layers were sintered by hot pressing. After post-processing the membranes were installed in a single cell test fixture. In addn., the ink used for coating was characterized in thin film form in sulfuric acid electrolyte and the results were compared to the fuel cell data. Current-voltage characteristics were measured after a running-in period of about 16 h. After cooling the test fixture down to room temp. the cyclic voltammograms were taken in situ by using a two-electrode potentiostatic measurement circuit. The electrode under study is purged by an inert gas (N₂) while the other electrode is purged by hydrogen gas so as to form a reversible hydrogen electrode. In this manner the surface processes taking place on both the anode and cathode side of the MEA could be measured and evaluated. From the hydrogen desorption charge on Pt the available Pt surface area has been detd. and the ratio between measured area and calcd. total Pt area was in the range of 30%. The ratios between electrochem. surface areas on the anode and cathode side compare well with the corresponding ratios obtained from calcns. of total added Pt surface area. The relation between the voltammetrically obtained surface areas and the polarization data (current-voltage curves) is discussed. An enlargement of the active surface area could be noted after potential scans into the Pt oxide formation region. This indicates the presence of a deactivation process, probably due to adsorbed org. material removed oxidatively during the extended polarization. When MEAs based on Pt/Ru alloy catalysts were measured the hydrogen adsorption charge decreased as a result of the lower hydrogen adsorption capability of Ru. In the double layer region the formation of Ru-oxides was also well manifested.

RE.CNT 3 THERE ARE 3 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 9 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 2001:73418 CAPLUS
DN 134:103343
TI Method for fabricating membrane and electrode assembly for polymer electrolyte membrane fuel cells
IN Kim, Chang-Soo; Chun, Young-Gap; Peck, Dong-Hyun; Shin, Dong Ryul
PA Korea Institute of Energy Research, S. Korea
SO U.S., 8 pp.
CODEN: USXXAM
DT Patent
LA English

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6180276	B1	20010130	US 1998-149088	19980908
PRAI	KR 1998-5531	A	19980223		

AB A method is disclosed for fabricating an electrode and membrane assembly (MEA) for polymer electrolyte membrane fuel cells.

The MEA comprises a polymer electrolyte membrane on each side of which an electrocatalyst layer is provided in a melted state. A perfluorosulfonyl fluoride copolymer powder ranging, in particle size distribution, from 20 to 200 μm is hot-pressed at 200-250.degree. to give a pre-formed sheet whose opposite sides are then coated with a catalyst **ink** consisting of Pt/C powder, glycerol and water. This catalyst **ink**-coated preformed sheet is again subjected to hot **pressing** at 200-250.degree. to embed the catalyst **ink** into the pre-formed sheet. Hydrolysis in NaOH/methanol or H₂SO₄ soln. converts the membrane of the sheet from a non-ionized form to an ionized form.

RE.CNT 2 THERE ARE 2 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 10 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 2001:320269 CAPLUS
DN 134:329056
TI Production of catalytic coatings on membranes suitable for low-temp. **fuel cells**
IN Gupta, Ashok Kumar; Tietz, Frank; Buchkremer, Hans Peter; Kundler, Isabel
PA Forschungszentrum Juelich G.m.b.H., Germany
SO PCT Int. Appl., 18 pp.
CODEN: PIXXD2
DT Patent
LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2001031725	A1	20010503	WO 2000-EP10129	20001014
	W: CA, JP, US				
	RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
	EP 1236233	A1	20020904	EP 2000-969506	20001014
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, FI, CY				

PRAI DE 1999-19951936 A 19991028
WO 2000-EP10129 W 20001014

AB A membrane, coated by a catalytic active coating, is suitable for low-temp. **fuel cells**. The prepn. of the membrane is carried out by several steps: (1) prodn. of a paste consisting of a non-polar solvent, catalytic active material, as well as a polymeric soln., (2) laminar coating, esp. by **silk screen printing**, of the paste on a polymer membrane with a thickness of 10-100 μm , (3) drying of the paste at 30-80.degree. and **pressing** with the membrane. The membrane coated with the paste is **pressed** at 100-150.degree., whereby the solvent is volatile and the catalytic active material is fixed on the membrane. A binder, or plasticizer is added to the paste. Nafion is used as polymer, platinum as catalyst, and terpineol as solvent. The paste comprises solids 10-50, preferably 20-30, binder 0-10, preferably 1-2, plasticizer 0-5, preferably 1-2 wt.%, and the rest is solvent. Optionally, the paste contains 60-80 wt.% solvent. The use of non-polar solvent prevents swelling of the membrane.

RE.CNT 5 THERE ARE 5 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 11 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 2002:69613 CAPLUS
DN 136:121086
TI Manufacture of electrode-electrolyte composites for **fuel cells** and substrates therefor
IN Inuzuka, Kyoko
PA Toyota Motor Corp., Japan

SO Jpn. Kokai Tokkyo Koho, 8 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 2002025578	A2	20020125	JP 2000-210739	20000712
PRAI	JP 2000-210739		20000712		

AB The composites are manufd. by applying an **ink** contg. catalysts on a porous substrate, **pressing** the substrate with a polymer electrolyte membrane, and sending a fluid (e.g., compressed air) to the substrate to remove the substrate from the catalyst **ink**. Preferably, the substrate has a surface roughness of water contact angle $\geq 120^\circ$. The process prevents damaging of the catalyst layer as the electrode.

L55 ANSWER 12 OF 20 CAPLUS COPYRIGHT 2003 ACS

AN 2002:810059 CAPLUS

DN 138:139989

TI Novel polymer electrolytes for microfabricated **fuel cells**

AU Wainright, J. S.; Litt, M. H.; Zhang, Y.; Liu, C. C.; Savinell, R. F.

CS Department of Chemical Engineering, Case Western Reserve University, Cleveland, OH, 44106-7217, USA

SO Proceedings - Electrochemical Society (2001), 2000-22 (Power Sources for the New Millennium), 14-22

CODEN: PESODO; ISSN: 0161-6374

PB Electrochemical Society

DT Journal

LA English

AB The objective of this research is to produce a hydrogen fueled, air-breathing **fuel cell** by combining microfabrication techniques with polymer electrolyte **fuel cell** technol.

Ideally, this will result in a low cost, easily manufd. device suitable for low power (μ W to mW) applications. It is envisioned that this device would operate in a passive mode, without active control of temp., humidity, reactant **pressure** or flow rate. The properties of the polymer electrolyte are crit. for successful operation under these conditions. One family of electrolytes under consideration are polyimide copolymers in which one of the monomers contains a sulfonic acid functionality to provide proton cond. and a nonionic monomer will provide dimensional stability. These electrolytes have enhanced cond. at low relative humidities and low permeability to hydrogen and oxygen, necessary requirements for a completely passive device.

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD

ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 13 OF 20 CAPLUS COPYRIGHT 2003 ACS

AN 2000:665700 CAPLUS

DN 133:225598

TI Method for applying electrode layers on a tape-like polymer electrolyte membrane for **fuel cells**

IN Starz, Karl-Anton; Zuber, Ralf; Gottenauer, Wolfgang; Fehl, Knut; Diehl, Manfred

PA Degussa-Huels Aktiengesellschaft, Germany

SO Eur. Pat. Appl., 15 pp.

CODEN: EPXXDW

DT Patent

LA German

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE

PI EP 1037295 A1 20000920 EP 1999-108791 19990503
EP 1037295 B1 20010613
R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT,
IE, SI, LT, LV, FI, RO
DE 19910773 A1 20000928 DE 1999-19910773 19990311
ES 2159978 T3 20011016 ES 1999-108791 19990503
CA 2300226 AA 20000911 CA 2000-2300226 20000309
JP 2000268829 A2 20000929 JP 2000-67017 20000310

PRAI DE 1999-19910773 A 19990311

AB In the title method, the front and back sides of the membrane (in the desired pattern) are continuously **pressed** with the electrode layers under application of an electrocatalyst-contg. **ink**, and the **pressed** electrode layers are dried at a high temp. immediately after the **press** process. By **pressing** under maintenance of a correct position arrangement of the pattern the electrode layers of front and back sides are formed to each other. X

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD
ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 14 OF 20 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-682608 [73] WPIDS
DNN N2002-538987 DNC C2002-192473
TI **Inks** for making anode and cathode catalysts for **ionomeric** membranes of direct methanol fuels, includes **platinum** and **platinum-ruthenium** catalysts, purified water, and perfluorosulfonic acid **ionomer**.
DC A85 G02 J04 L03 X16
IN DAVEY, J; GOTTESFELD, S; REN, X; THOMAS, S C; ZELENAY, P
PA (REGC) UNIV CALIFORNIA
CYC 94
PI WO 2002045188 A2 20020606 (200273)* EN 20p
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ
LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD
SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW
AU 2002039439 A 20020611 (200273)

ADT WO 2002045188 A2 WO 2001-US45619 20011031; AU 2002039439 A AU 2002-39439
20011031

FDT AU 2002039439 A Based on WO 200245188

PRAI US 2000-715211 20001114

AB WO 200245188 A UPAB: 20021113

NOVELTY - **Inks** comprise a **platinum** catalyst for a cathode catalyst and a **platinum-ruthenium** catalyst for an anode catalyst, purified water which is 4-20 times that of the catalyst by weight, and a perfluorosulfonic acid **ionomer** in an amount effective to provide an **ionomer** content in dried anode and cathode catalyst coats of 20-80 vol.%. X

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for the following:

(a) the production of **inks** for use in anode and cathode catalysts applied to membranes for direct methanol **fuel cells**, comprising combining **platinum** or **platinum-ruthenium** catalysts with purified water to form a first mixture, cooling the first mixture to a temperature that reduces evaporation of water from the first mixture, placing the first mixture in a cooler and agitating the mixture to obtain a first homogeneous suspension, adding an alcoholic solution of perfluorosulfonic acid to the first homogeneous suspension to provide a second mixture, agitating the second mixture in the cooler to obtain a homogeneous **ink** suitable for application to the appropriate anode or cathode side of the membrane; and

(b) a method for applying anode and cathode catalyst to anode and

cathode surfaces for a direct methanol fuel cell, comprising forming a mixture of water, perfluorosulfonic acid ionomer, alcohols, and catalyst of platinum-ruthenium for the anode or platinum catalyst for the cathode while cooling and agitating the solution, placing the mixture in a cooler, continuously agitating the mixture, spraying the mixture on the anode or cathode surfaces as determined by the catalyst content.

USE - For making anode and cathode catalysts for application to ionomer membranes for direct methanol fuels.

Dwg.0/5

L55 ANSWER 15 OF 20 WPIDS (C) 2003 THOMSON DERWENT
AN 2000-563400 [52] WPIDS
DNN N2001-155876 DNC C2001-065378
TI Fabrication of membrane and electrode assembly by coating an electrocatalyst ink of carbon-supported platinum, glycerol and water on a pre-formed copolymer sheet, and embedding the electrocatalyst coats by hot **pressing**.
DC A85 L03 X16
IN CHUN, Y G; KIM, C S; PECK, D H; SHIN, D R; BAEK, D H; JEON, Y G; SHIN, D Y; CHUN, Y; KIM, C; PECK, D
PA (KOEN-N) KOREA INST ENERGY RES
CYC 2
PI KR 99070596 A 19990915 (200052)*
US 6180276 B1 20010130 (200123)B 8p
KR 263992 B1 20000816 (200134)
ADT KR 99070596 A KR 1998-5531 19980223; US 6180276 B1 US 1998-149088 19980908; KR 263992 B1 KR 1998-5531 19980223
PRAI KR 1998-5531 19980223
AB US 6180276 B UPAB: 20010425 ABEQ treated as Basic
NOVELTY - A membrane and electrode assembly is fabricated by coating an electrocatalyst ink on a pre-formed perfluorosulfonyl fluoride copolymer sheet, embedding electrocatalyst coats into the sheet by hot **pressing** the electrocatalyst-coated sheet, and hydrolyzing the electrocatalyst-embedded sheet. The catalyst ink comprises carbon-supported platinum, glycerol, and water. X

DETAILED DESCRIPTION - Fabrication of membrane and electrode assembly (MEA) involves pre-forming a perfluorosulfonyl fluoride copolymer powder into a sheet by hot **pressing** process at 200-250 deg. C, coating an electrocatalyst ink on either side of pre-formed sheet by **screen printing** process and drying the coats, embedding the electrocatalyst coats into the sheet by hot **pressing** the electrocatalyst-coated sheet, and hydrolyzing the electrocatalyst-embedded sheet to convert the membrane of the sheet from a non-ionized form into an ionized form. The ink comprises carbon-supported platinum (Pt/C), glycerol, and water in the absence of perfluorinated sulfonyl fluoride binder. The copolymer powder has 20-200 μ m particle size distribution.

USE - The method is used for fabricating membrane and electrode assembly for polymer electrolyte membrane fuel cells (PEMFC). The PEMFC is used as an electric power source for, e.g. automobiles, as transportable electric source or an on-site power source.

ADVANTAGE - The inventive method is simpler and economically much more favorable than the conventional methods which use **Nafion** solution. The MEA fabricated by the method is far superior to conventional MEAs in the bonding strength between electrode and membrane and in the effective electrode reaction on the three-phase boundary.

Dwg.0/7

AB KR 99070596 A UPAB: 20010502
NOVELTY - A membrane and electrode assembly is fabricated by coating an electrocatalyst ink on a pre-formed perfluorosulfonyl fluoride copolymer sheet, embedding electrocatalyst coats into the sheet by hot **pressing** the electrocatalyst-coated sheet, and hydrolyzing the

electrocatalyst-embedded sheet. The catalyst **ink** comprises carbon-supported **platinum**, glycerol, and water.

DETAILED DESCRIPTION - Fabrication of membrane and electrode assembly (MEA) involves pre-forming a perfluorosulfonyl fluoride copolymer powder into a sheet by hot **pressing** process at 200-250 deg. C, coating an electrocatalyst **ink** on either side of pre-formed sheet by **screen printing** process and drying the coats, embedding the electrocatalyst coats into the sheet by hot **pressing** the electrocatalyst-coated sheet, and hydrolyzing the electrocatalyst-embedded sheet to convert the membrane of the sheet from a non-ionized form into an ionized form. The **ink** comprises carbon-supported **platinum (Pt/C)**, glycerol, and water in the absence of perfluorinated sulfonyl fluoride binder. The copolymer powder has 20-200 μm particle size distribution.

USE - The method is used for fabricating membrane and electrode assembly for polymer electrolyte membrane **fuel cells** (PEMFC). The PEMFC is used as an electric power source for, e.g. automobiles, as transportable electric source or an on-site power source.

ADVANTAGE - The inventive method is simpler and economically much more favorable than the conventional methods which use **Nafion** solution. The MEA fabricated by the method is far superior to conventional MEAs in the bonding strength between electrode and membrane and in the effective electrode reaction on the three-phase boundary.

Dwg. 0/7

L55 ANSWER 16 OF 20 CAPLUS COPYRIGHT 2003 ACS
AN 1998:805108 CAPLUS
DN 130:116423
TI Current efficiency for soybean oil hydrogenation in a solid polymer electrolyte reactor
AU An, W.; Hong, J.-K.; Pintauro, P. N.
CS Department of Chemical Engineering, Tulane University, New Orleans, LA, 70118, USA
SO Journal of Applied Electrochemistry (1998), 28(9), 947-954
CODEN: JAELBJ; ISSN: 0021-891X
PB Chapman & Hall
DT Journal
LA English
AB Soybean oil was hydrogenated electrocatalytically in a solid polymer electrolyte (SPE) reactor, similar to that in H₂/O₂ **fuel cells**, with water as the anode feed and source of hydrogen. The key component of the reactor was a membrane electrode assembly (MEA), composed of a precious metal-black cathode, a RuO₂ powder anode, and a **Nafion** 117 cation-exchange membrane. The SPE reactor was operated in a batch recycle mode at 60.degree. and one atm. **pressure** using a com.-grade soybean oil as the cathode feed. Various factors that might affect the oil hydrogenation current efficiency were studied, including the type of cathode catalyst, catalyst loading, the cathode catalyst binder loading, c.d., and reactant flow rate. The current efficiency ordering of different cathode catalysts is Pd > **Pt** > Rh > Ru > Ir. Oil hydrogenation current efficiencies with a Pd-black cathode decreased with increasing c.d. and ranged from .apprx.70% at 0.050 A cm⁻² to 25% at 0.490 A cm⁻². Current pulsing for frequencies at 0.25-60 Hz had no effect on current efficiencies. The optimum cathode catalyst loading for both Pd and **Pt** was 2.0 mg cm⁻². Soybean oil hydrogenation current efficiencies was unaffected by **Nafion** and PTFE cathode catalyst binders, as long as the total binder content was .1toreq.30% (based on the dry catalyst wt.). When the oil feed flow rate was increased from 80 to 300 mL min⁻¹, the oil hydrogenation current efficiency at 0.100 A cm⁻² increased from 60% to 70%. A high (70%) current efficiency was achieved at 80 mL min⁻¹ by inserting a nickel **screen turbulence promoter** into the oil stream.

ALL CITATIONS AVAILABLE IN THE RE FORMAT

L55 ANSWER 17 OF 20 WPIDS (C) 2003 THOMSON DERWENT
 AN 1996-022133 [03] WPIDS
 DNN N1996-018387 DNC C1996-007681
 TI Catalytic gas diffusion electrode for a **fuel cell** with solid electrolyte - has hydrophobic porous back support, intermediate layer of electron and proton conductor materials, and a catalytic active layer contg. catalyst and proton conductor.
 DC A85 L03 X16
 IN DIRVEN, P; ENGELEN, W
 PA (VITO-N) VITO VLAAMSE INSTELLING TECHNOLOGISCH
 CYC 19
 PI EP 687023 A1 19951213 (199603)* EN 8p
 R: AT CH DE DK ES FR GB GR IE IT LI LU NL PT SE
 WO 9534098 A1 19951214 (199604) EN 21p
 W: JP
 CA 2151104 A 19951208 (199614)
 BE 1008455 A3 19960507 (199625) NL 22p
 US 5561000 A 19961001 (199645) 6p
 JP 09501541 W 19970210 (199716) 17p
 EP 687023 B1 19980304 (199813) EN 9p
 R: AT CH DE DK ES FR GB GR IE IT LI LU NL PT SE
 DE 69501681 E 19980409 (199820)
 ADT EP 687023 A1 EP 1995-201463 19950602; WO 9534098 A1 WO 1995-BE53 19950607;
 CA 2151104 A CA 1995-2151104 19950606; BE 1008455 A3 BE 1994-561 19940607;
 US 5561000 A US 1995-465110 19950605; JP 09501541 W WO 1995-BE53 19950607,
 JP 1996-500096 19950607; EP 687023 B1 EP 1995-201463 19950602; DE 69501681
 E DE 1995-601681 19950602, EP 1995-201463 19950602
 FDT JP 09501541 W Based on WO 9534098; DE 69501681 E Based on EP 687023
 PRAI BE 1994-561 19940607
 AB EP 687023 A UPAB: 19960122
 Catalytic gas diffusion electrode for a solid electrolyte cell, esp. a **fuel cell**, has a hydrophobic porous back support, a non-catalytic intermediate layer contg. a mixt. of electron-conductive material (I) and proton-conductive **ionomer** (II), and a catalytic active layer in which the catalyst particles are bound by (II).
 Prodn. of the electrode involves: depositing the intermediate layer in the form of an **ink** contg. (I) and (II); and forming the active layer in the form of an **ink** contg. catalyst and (II).
 Catalyst is platinised C. Intermediate layer is formed of C powder as (I) and **NAFION** (RTM) as (II). Back layer is a C cloth filled with C powder and binder, pref. PTFE.
 Back layer is pref. formed by pasting the C cloth with a blade. The intermediate and catalytic layers are formed by applying the **ink** in multiple stages with intermediate drying steps.
 In an example, the back support is formed of a C cloth filled with a paste of 66% C and 34% PTFE emulsion. The **ink** for the intermediate layer contains 10g C, 11 5% **NAFION** and 11 ethanol. The **ink** for the active layer contains 50g platinised C, 11 5% **NAFION** and 11 ethanol. The dried intermediate layer has a thickness of a few micron and the Pt loading of the outer active layer is 0.43 mg/sq.cm..
 USE - Esp. as the O₂ electrode of a **fuel cell**.
 ADVANTAGE - Electrode promotes highly efficient use of the catalyst and can work with a large output under air at atmos. pressure.
 Dwg. 0/2

L55 ANSWER 18 OF 20 WPIDS (C) 2003 THOMSON DERWENT
 AN 1993-175426 [21] WPIDS
 CR 1992-316385 [38]
 DNN N1993-134487 DNC C1993-078368
 TI Solid polymer electrolyte membrane mfr. for **fuel cell**

- by forming film as **ink** that is spread and cured on film
release blank the transferring film to membrane and hot **pressing**

DC A85 L03 X16
IN WILSON, M S
PA (REGC) UNIV CALIFORNIA
CYC 1
PI US 5211984 A 19930518 (199321)* 11p ✓
ADT US 5211984 A CIP of US 1991-656329 19910219, CIP of US 1991-736876
19910729, US 1991-811220 19911220
PRAI US 1991-656329 19910219; US 1991-736876 19910729; US 1991-811220
19911220
AB US 5211984 A UPAB: 19931116
The mfr. of a SPE membrane assembly comprises (a) furnishing a SPE
membrane in Na(+) form; (b) furnishing a perfluorosulphonate
ionomer in a Na(+) form or thermoplastic form; (c) uniformly
dispersing a supported Pt catalyst and a solvent in the Na(+) or
thermoplastic form of the **ionomer** for form an **ink**; (d)
forming a film of the **ink** contg. a predetermined amt. of the
catalyst on a surface of the SPE membrane in the Na(+) form; (e) heating
the film of the **ink** to a temp. effective to dry the **ink**
; and (f) converting the film of the **ink** and the SPE membrane to
a protonated form of perfluorosulphonate.

USE/ADVANTAGE - Mfr. of an solid polymer electrolyte (SPE) membrane
is provided. It is useful for gas **fuel cells**. The
catalyst layer has adequate gas permeability so that the cell performance
is not affected and has a density and particle distribution effective to
optimise proton access to the catalyst and electronic continuity for
electron flow from the half-cell reaction occurring at the catalyst. SPE
fuel cells are provided with relatively low supported
catalyst loadings with no redn. in their performance. The bonding between
the SPE layer and the catalyst layer is improved. The wt. fraction of the
SPE **ionomer** with the catalyst layer is increased to improve the
efficiency of the catalyst layer.

Dwg.1/7

L55 ANSWER 19 OF 20 WPIDS (C) 2003 THOMSON DERWENT
AN 2000-025283 [03] WPIDS
DNN N2000-018945 DNC C2000-006486
TI Production of slurry for proton exchange membrane **fuel**
cell catalyst layer.
DC A85 L03 X16
IN CHANG, H; LIM, C
PA (SMSU) SAMSUNG ELECTRONICS CO LTD
CYC 27
PI EP 955687 A2 19991110 (200003)* EN 12p
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI
KR 99087023 A 19991215 (200056)
US 6344428 B1 20020205 (200211) X
ADT EP 955687 A2 EP 1999-300520 19990125; KR 99087023 A KR 1998-60009
19981229; US 6344428 B1 US 1999-258802 19990226
PRAI KR 1998-60009 19981229; KR 1998-15991 19980504
AB EP 955687 A UPAB: 20000118
NOVELTY - A slurry for the catalyst layer of a proton exchange membrane
fuel cell is made by forming an alkali metal salt of a
perfluorosulfonate **ionomer** solution, adding a polar organic
solvent, heating to remove alcohol and adding **platinum**/carbon.
DETAILED DESCRIPTION - A method of making a slurry for forming a
catalyst layer of a proton exchange membrane (PEM) **fuel**
cell comprises adding an MOH solution (M = Li, Na or K) to a
perfluorosulfonate **ionomer** (PFSI) solution to give an M+-PFSI
solution, adding an organic polar solvent of higher boiling point than any

alcohol remaining in the PFSI solution and heating to between the alcohol boiling point and 20 deg. C above this to remove remaining alcohol. The pretreated PFSI solution is then mixed with Pt/C to form the slurry. An INDEPENDENT CLAIM is also included for a method as above of making a PEM fuel cell further comprising coating the slurry on one side of an electrode backing layer, drying at not more than the boiling point of the polar solvent, impregnating in acid solution, washing and drying to form a gas diffusion electrode with the deposited catalyst layer and interposing a PEM between anode and cathode side of the electrode and hot pressing.

USE - In forming a catalyst layer for a PEM fuel cell electrode (claimed).

ADVANTAGE - The electrode has improved processing stability and power output characteristics. Continuous production by tape casting is possible and only one coating step is needed.

DESCRIPTION OF DRAWING(S) - A block diagram of the process steps is shown.

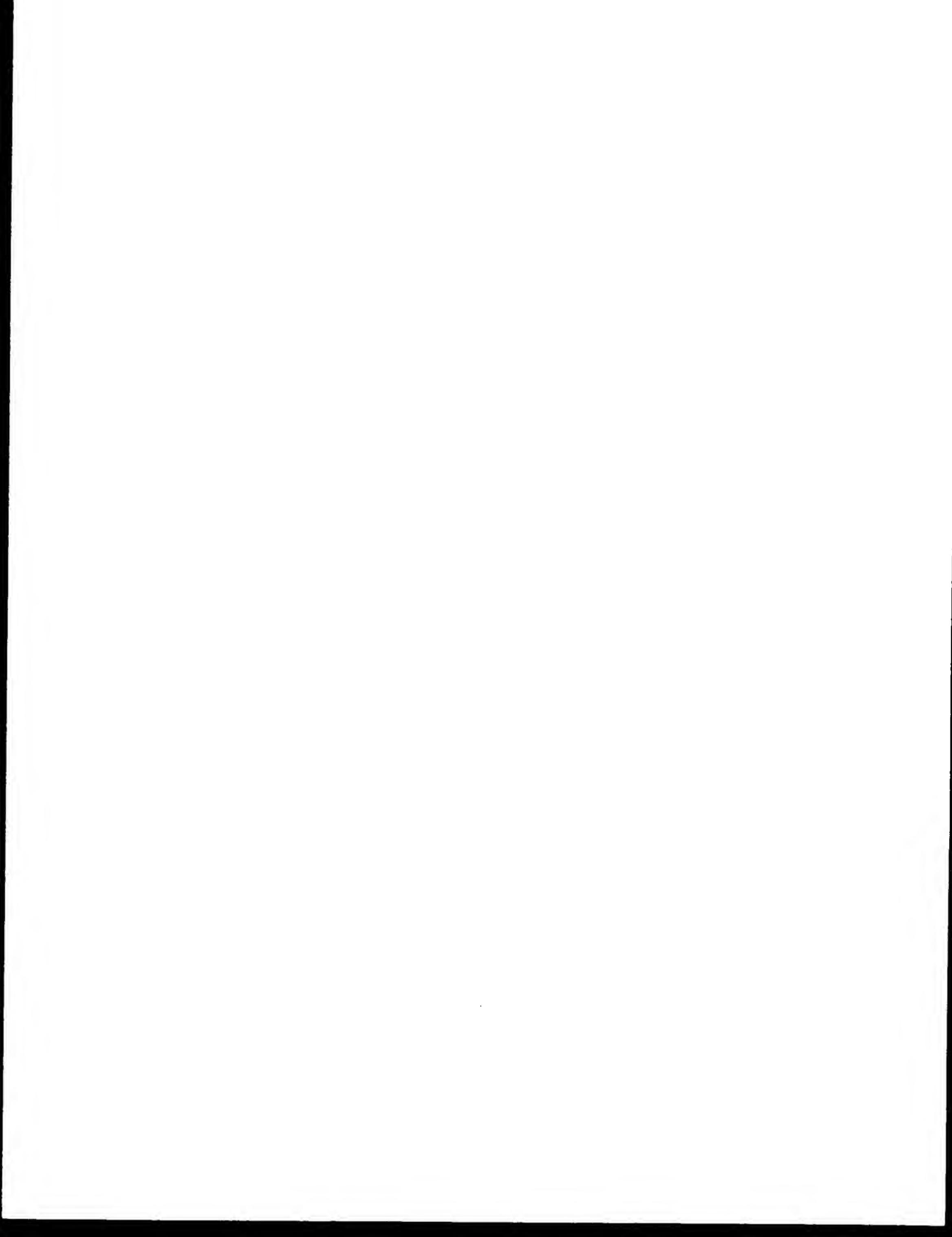
Dwg.2/6

L55 ANSWER 20 OF 20 WPIDS (C) 2003 THOMSON DERWENT
AN 2003-054401 [05] WPIDS
DNC C2003-014000
TI Self-humidifying polymer membrane for polymer fuel cell and process for producing self-humidifying polymer membrane/electrode laminate by using the same.
DC A85 L03 X16
IN BAEK, D H; KIM, C S; KWAK, S H; PARK, G G; YANG, T H; YOON, G H
PA (KOEN-N) KOREA INST ENERGY RES
CYC 1
PI KR 2002030963 A 20020426 (200305)* 1p
KR 343117 B 20020705 (200305)
ADT KR 2002030963 A KR 2000-61496 20001019; KR 343117 B KR 2000-61496 20001019
FDT KR 343117 B Previous Publ. KR 2002030963
PRAI KR 2000-61496 20001019
AB KR2002030963 A UPAB: 20030121

NOVELTY - Provided are a self-humidifying polymer membrane for a polymer fuel cell, which is produced by using a sputtering method, and a process for producing a self-humidifying polymer membrane/electrode laminate by using the self-humidifying polymer membrane.

DETAILED DESCRIPTION - The self-humidifying polymer membrane is produced by a process comprising the steps of: forming a pre-foamed precursor sheet by hot pressing or rolling a perfluorosulfonyl fluoride/TFE copolymer resin; laminating platinum particles on one side of the pre-foamed precursor sheet by sputtering; laminating closely other pre-foamed precursor sheet on the platinum laminated side; making a pre-foamed precursor sheet laminate by hot pressing or rolling the laminated two pre-foamed precursor sheets; impregnating the pre-foamed precursor sheet laminate in a caustic soda solution to produce Na⁺ type perfluoro sulfonate polymer membrane. And the process for producing the self-humidifying polymer membrane/electrode laminate comprises the steps of: making the electrode by coating a film with a catalyst ink comprising platinum coated carbon, a nafion solution, glycerol, and tetrabutyl ammonium hydroxide; laminating the electrode on the Na⁺ type polymer membrane and hot pressing; soaking the self-humidifying polymer membrane/electrode laminate in a sulfuric acid solution to produce H⁺ type polymer membrane/electrode laminate; cleansing the H⁺ type polymer membrane/electrode laminate with deionized pure water many times.

Dwg.1/10



M. S.
Wilson

L68 ANSWER 1 OF 7 WPIDS (C) 2003 THOMSON DERWENT
AN 2000-256445 [22] WPIDS
DNN N2000-190703 DNC C2000-078188
TI Fuel cell system includes a fuel cell stack comprising fuel cells having membrane/electrode assemblies that are hydrated with liquid water, and bipolar plates for distributing hydrogen fuel gas and water.
DC L03 X16
IN WILSON, M S
PA (REGC) UNIV CALIFORNIA
CYC 85
PI WO 2000011745 A1 20000302 (200022)* EN 28p H01M008-04
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL
OA PT SD SE SL SZ UG ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB
GE GH GM HR HU ID IL IS JP KE KG KP KR KZ LC LK LR LS LT LU LV MD
MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR TT UA
UG UZ VN YU ZA ZW
AU 9957717 A 20000314 (200031)
US 6117577 A 20000912 (200046)
EP 1110264 A1 20010627 (200137) EN H01M008-04
R: AL AT BE CH CY DE DK ES FI FR GB GR IE IT LI LT LU LV MC MK NL PT
RO SE SI
EP 1110264 B1 20030402 (200325) EN H01M008-04
R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE
ADT WO 2000011745 A1 WO 1999-US17573 19990803; AU 9957717 A AU 1999-57717
19990803; US 6117577 A US 1998-135965 19980818; EP 1110264 A1 EP
1999-945013 19990803, WO 1999-US17573 19990803; EP 1110264 B1 EP
1999-945013 19990803, WO 1999-US17573 19990803
FDT AU 9957717 A Based on WO 200011745; EP 1110264 A1 Based on WO 200011745;
EP 1110264 B1 Based on WO 200011745
PRAI US 1998-135965 19980818
IC ICM H01M008-04
ICS H01M008-02
AB WO 200011745 A UPAB: 20000508
NOVELTY - Ambient pressure fuel cell system includes a fuel stack comprising fuel cells having membrane/electrode assemblies (MEA's) (24) that are hydrated with liquid water, and bipolar plates (26) with anode and cathode channels for respectively distributing hydrogen fuel gas and water to the anode side and air with reactant oxygen to a cathode side.
DETAILED DESCRIPTION - Ambient pressure fuel cell system includes a fuel stack comprising of fuel cells having membrane/electrode assemblies (MEA's) that are hydrated with liquid water, and bipolar plates with anode and cathode channels for respectively distributing hydrogen fuel gas and water to the anode side and air with reactant oxygen to a cathode side. The system also includes a liquid water supply to the fuel cells for hydrating the MEA's, a hydrogen fuel gas supply, and near-ambient pressure blower for blowing air in excess of reaction stoichiometric amounts through the fuel cell stack to provide oxygen for electrochemical reaction at the cathode side.
An INDEPENDENT CLAIM is also included for a method of operating a fuel cell stack at ambient pressure.
USE - The system of this invention is useful as hydrogen-oxygen fuel cells, and more particularly, to polymer electrolyte membrane fuel cells.
ADVANTAGE - The system of this invention has fully hydrated membranes that overcomes the problems inherent in pressurized fuel cells using humidified reactant gases. The system of this invention uses direct liquid hydration at ambient pressure to

produce a simple, low-parasitic power system and to provide direct evaporative-cooling from high volume ambient **pressure** air flow.

DESCRIPTION OF DRAWING(S) - Figure showing a cross-section of a unit cell defining an ambient **fuel cell** system of this invention.

Flow field 12

Membrane/electrode assemblies 24

Bipolar plates 26

Dwg. 1/7

FS CPI EPI

FA AB; GI

MC CPI: L03-E04

EPI: X16-C01C; X16-E06A

L68 ANSWER 2 OF 7 CAPLUS COPYRIGHT 2003 ACS

AN 1992:87618 CAPLUS

DN 116:87618

TI Thin-film catalyst layers for polymer electrolyte **fuel cell** electrodes

AU Wilson, M. S.; Gottesfeld, S.

CS Electron. Res. Group, Los Alamos Natl. Lab., Los Alamos, NM, 87545, USA

SO Journal of Applied Electrochemistry (1992), 22(1), 1-7

CODEN: JAELBJ; ISSN: 0021-891X

DT Journal

LA English

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)

Section cross-reference(s): 38, 72

AB New structures for the Pt/C catalyst layer of polymer electrolyte **fuel cell** electrodes have been developed that substantially increase the utilization efficiency of the catalyst. Fabricating the catalyst layers and gas diffusion backings sep. makes it possible to formulate each structure with the properties that are most suitable for its function. In the case of the catalyst layer, the optimal properties are hydrophilicity, thinness, uniformity, and the proper ratio of ionomer and supported catalyst. The catalyst layers are cast from soln. as thin films that utilize the ionomer itself as a binder. The thin films are hot **pressed** directly onto the ionomer membranes, and the hydrophobic gas diffusion backings are inserted when the cells are assembled. The performances of **fuel cells** based on the thin film catalyst layers are comparable with those of gas diffusion electrode designs that utilize several times as much Pt, thus the specific activities of the Pt catalyst in the new structures are significantly higher.

ST polymer electrolyte **fuel cell** electrode; platinum carbon catalyst electrode **fuel cell**

IT Carbon black, uses

RL: USES (Uses)

(electrodes, thin-film platinum catalytic, for polymer electrolyte **fuel cell**)

IT Electrolytic polarization

(of platinum-carbon thin film cathodes in **fuel cell** with Nafion 117 membrane)

IT Polyoxyalkylenes, uses

RL: USES (Uses)

(fluorine- and sulfo-contg., ionomers, binder, in fabrication of platinum-carbon thin film cathodes for **fuel cells**)

IT Electrodes

(**fuel-cell**, catalytic, platinum- carbon thin film layer for, fabrication and performance of, for polymer electrolyte **fuel cells**)

IT Fluoropolymers

RL: USES (Uses)

(polyoxyalkylene-, sulfo-contg., ionomers, binder, in fabrication of

Hand

havre

platinum-carbon thin film cathodes for fuel cells)

IT Ionomers
 RL: USES (Uses)
 (polyoxyalkylenes, fluorine- and sulfo-contg., binder, in fabrication
 of platinum-carbon thin film cathodes for fuel cells
)

IT 7440-06-4, Platinum, uses
 RL: USES (Uses)
 (catalyst electrodes, thin-film, for polymer electrolyte fuel
 cell)

IT 56-81-5, Glycerol, uses
 RL: USES (Uses)
 (inks contg., in fabrication of thin-film catalyst layers for polymer
 electrolyte fuel cell electrodes)

IT 66796-30-3, Nafion 117
 RL: USES (Uses)
 (membranes, fuel cell with thin film
 platinum-carbon cathodes and, polarization of)

RN 7440-06-4

RN 56-81-5

RN 66796-30-3

L68 ANSWER 3 OF 7 WPIDS (C) 2003 THOMSON DERWENT

AN 1992-316385 [38] WPIDS

CR 1993-175426 [21]

DNN N1992-242081 DNC C1992-140580

TI Gas reaction fuel cell having solid polymer
 electrolyte - has film of platinum catalyst in ionomer between solid
 polymer electrolyte and cathode.

DC A85 J04 L03 X16

IN WILSON, M S

PA (REGC) UNIV CALIFORNIA LOS ALAMOS NAT LAB; (LOSA-N) LOS ALAMOS NAT LAN;
 (REGC) UNIV CALIFORNIA; (USAT) US DEPT ENERGY; (LALA-N) LOS ALAMOS NAT LAB

CYC 17

PI WO 9215121 A1 19920903 (199238)* EN 29p H01M008-10
 RW: AT BE CH DE DK ES FR GB GR IT LU MC NL SE
 W: CA JP

US 656329 A0 19930501 (199322) 21p H01M000-00

US 5234777 A 19930810 (199333) 11p H01M008-10

JP 05507583 W 19931028 (199348) 8p H01M004-92

EP 600888 A1 19940615 (199423) EN H01M008-10
 R: DE FR GB

EP 600888 B1 19970827 (199739) EN 14p H01M008-10
 R: DE FR GB

DE 69221881 E 19971002 (199745) H01M008-10

ADT WO 9215121 A1 WO 1992-US1058 19920218; US 656329 A0 US 1991-656329
 19910219; US 5234777 A CIP of US 1991-656329 19910219, US 1991-736876
 19910729; JP 05507583 W JP 1992-507039 19920218, WO 1992-US1058 19920218;
 EP 600888 A1 EP 1992-907218 19920218, WO 1992-US1058 19920218; EP 600888
 B1 EP 1992-907218 19920218, WO 1992-US1058 19920218; DE 69221881 E DE
 1992-621881 19920218, EP 1992-907218 19920218, WO 1992-US1058 19920218

FDT JP 05507583 W Based on WO 9215121; EP 600888 A1 Based on WO 9215121; EP
 600888 B1 Based on WO 9215121; DE 69221881 E Based on EP 600888, Based on
 WO 9215121

PRAI US 1991-736876 19910729; US 1991-656329 19910219

REP 2.Jnl.Ref; US 4215183; US 4262063; US 4369103; US 4804592; US 4876115;
 7.Jnl.Ref; US 3134697

IC ICM H01M000-00; H01M004-92; H01M008-10

ICS H01M008-02

AB WO 9215121 A UPAB: 20011211
 A gas reaction fuel cell (10) has a solid polymer
 electrolyte (26) sepg. anode (16) and cathode (18) electrodes. A composite
 film (22) of a supported Pt catalyst (24) uniformly dispersed in a proton

conducting ionomer (28) is disposed between the solid polymer electrolyte (26) and the cathode (18). The platinum loading in the composite film (22) is less than 0.35 mg platinum per sq. cm. The composite film (22) is less than 10 microns thick.

USE/ADVANTAGE - The cell (10) is a H₂ O₂ fuel cell needing a catalyst (24) to proceed at useful rates. The composite film (22) optimises proton access to the catalyst (24) while giving adequate gas permeability and electronic continuity for electron flow from the half-cell reaction occurring at the catalyst.

Dwg.1/7

FS CPI EPI
FA AB; GI
MC EPI: X16-C01; X16-F02

L68 ANSWER 4 OF 7 WPIDS (C) 2003 THOMSON DERWENT
AN 1997-340925 [31] WPIDS
DNN N1997-282930 DNC C1997-109456
TI Polymer electrolyte membrane fuel cell - has macroporous flow-field with interdigitated inlet and outlet reactant channels contacting gas diffusion layer.
DC A85 L03 X16
IN WILSON, M S
PA (REGC) UNIV CALIFORNIA OFFICE TECHNOLOGY
CYC 1
PI US 5641586 A 19970624 (199731)* 12p H01M008-10
ADT US 5641586 A US 1995-568088 19951206
PRAI US 1995-568088 19951206
IC ICM H01M008-10
AB US 5641586 A UPAB: 19970731
A polymer electrolyte membrane fuel cell has: (a) a gas diffusion layer (26) with a first side contacting a catalytic surface (28) of the membrane (32); (b) a macro-porous flow-field layer (24) contacting the second side of the gas diffusion layer, for distributing a gaseous reactant over the layer for transport to the membrane; and (c) a reactant distribution plate, having interdigitated flow channels (22) contacting the flow-field layer, for delivering reactant to and removing reactant from the flow-field layer. Also claimed is a polymer electrolyte membrane fuel cell as above where the flow field layer defines the interdigitated flow channels. Further claimed is the field flow layer.

ADVANTAGE - the flow field has reduced pressure drop through the field. Accumulation of water in the flow field is minimised and access to the catalyst layers is maximised. Backings are very thin, to minimise the gas diffusional barrier.

Dwg.1B/7

FS CPI EPI
FA AB; GI
MC CPI: A12-E06; L03-E04
EPI: X16-C01

L68 ANSWER 5 OF 7 COMPENDEX COPYRIGHT 2003 EEI
AN 1992(10):123814 COMPENDEX DN 9210126227
TI Thin-film catalyst layers for polymer electrolyte fuel cell electrodes.
AU Wilson, M.S. (Los Alamos Natl Lab, Los Alamos, NM, USA); Gottesfeld, S.
SO J Appl Electrochem v 22 n 1 Jan 1992 p 1-7
CODEN: JAELBJ ISSN: 0021-891X
PY 1992
DT Journal
TC Experimental
LA English

haw

AB New structures for the Pt/C catalyst layer of polymer electrolyte fuel cell electrodes have been developed that substantially increase the utilization efficiency of the catalyst. Fabricating the catalyst layers and gas diffusion backings separately makes it possible to formulate each structure with the properties that are most suitable for its function. In the case of the catalyst layer, the optimal properties are hydrophilicity, thinness, uniformity, and the proper ratio of ionomer itself as a binder. The thin films are hot pressed directly onto the ionomer membranes, and the hydrophobic gas diffusion backings are inserted when the cells are assembled. The performances of fuel cells based on the thin film catalyst layers are comparable with those of gas diffusion electrode designs that utilize several times as much platinum, thus the specific activities of the Pt catalysts in the new structures are significantly higher. (Author abstract) 13 Refs.

CC 702 Electric Batteries & Fuel Cells; 802 Chemical Apparatus & Plants; 803 Chemical Agents & Basic Industrial Chemicals; 804 Chemical Products

CT *FUEL CELLS: Electrodes; CATALYSTS: Platinum; ELECTRODES, ELECTROCHEMICAL: Catalysis; IONOMERS; PLATINUM AND ALLOYS: Thin Films; ELECTROLYTES, SOLID

ST POLYMER ELECTROLYTE FUEL CELLS; THIN FILM CATALYSTS

ET Pt

L68 ANSWER 6 OF 7 WPIDS (C) 2003 THOMSON DERWENT
AN 1993-175426 [21] WPIDS
CR 1992-316385 [38]
DNN N1993-134487 DNC C1993-078368
TI Solid polymer electrolyte membrane mfr. for fuel cell
- by forming film as ink that is spread and cured on film release blank
the transferring film to membrane and hot pressing.
DC A85 L03 X16
IN WILSON, M S
PA (REGC) UNIV CALIFORNIA
CYC 1
PI US 5211984 A 19930518 (199321)* 11p H01M004-88
ADT US 5211984 A CIP of US 1991-656329 19910219, CIP of US 1991-736876
19910729, US 1991-811220 19911220
PRAI US 1991-656329 19910219; US 1991-736876 19910729; US 1991-811220
19911220
IC ICM H01M004-88
AB US 5211984 A UPAB: 19931116
The mfr. of a SPE membrane assembly comprises (a) furnishing a SPE membrane in Na(+) form; (b) furnishing a perfluorosulphonate ionomer in a Na(+) form or thermoplastic form; (c) uniformly dispersing a supported Pt catalyst and a solvent in the Na(+) or thermoplastic form of the ionomer for form an ink; (d) forming a film of the ink contg. a predetermined amt. of the catalyst on a surface of the SPE membrane in the Na(+) form; (e) heating the film of the ink to a temp. effective to dry the ink; and (f) converting the film of the ink and the SPE membrane to a protonated form of perfluorosulphonate.

USE/ADVANTAGE - Mfr. of an solid polymer electrolyte (SPE) membrane is provided. It is useful for gas fuel cells. The catalyst layer has adequate gas permeability so that the cell performance is not affected and has a density and particle distribution effective to optimise proton access to the catalyst and electronic continuity for electron flow from the half-cell reaction occurring at the catalyst. SPE fuel cells are provided with relatively low supported catalyst loadings with no redn. in their performance. The bonding between the SPE layer and the catalyst layer is improved. The wt. fraction of the SPE ionomer with the catalyst layer is increased to improve the efficiency of the catalyst layer.

Dwg.1/7

FS CPI EPI

FA AB; GI
MC CPI: A04-A; A04-E10D; A12-E06; A12-M02; L03-E04
EPI: X16-C01; X16-F02

L68 ANSWER 7 OF 7 WPIDS (C) 2003 THOMSON DERWENT
AN 2001-528505 [58] WPIDS

DNN N2001-392173

TI Bipolar plate for electrochemical cell e.g. **fuel cells**, uses electrically conductive foils on both sides of plate which make contact mutually at the openings at intersections of channels on both sides of support plate.

DC X16

IN WILSON, M S; ZAWODZINSKI, C
PA (REGC) UNIV CALIFORNIA

CYC 1

PI US 6255012 B1 20010703 (200158)* 8p H01M002-14

ADT US 6255012 B1 US 1999-444216 19991119

PRAI US 1999-444216 19991119

IC ICM H01M002-14

AB US 6255012 B UPAB: 20011010

NOVELTY - A bipolar plate (50) has multiple flow channels (52,54) formed on both sides of support plate (55), so that the channels on one side are perpendicular to channels on other side. The channels have a depth effective to form openings through the plate at their intersections (56). Electrically conductive foils (58,62) are provided on both sides of the plate which make electrical contact (64) mutually at the openings.

DETAILED DESCRIPTION - An INDEPENDENT CLAIM is also included for ~~fuel cell assembly.~~

USE - In **fuel cell** assembly (claimed) to separate fuel and oxidizing reaction gases on opposite sides of plate.

ADVANTAGE - Withstands high clamping **pressure** with the use of metal foils that are thin enough to be readily stamped. Light in weight and inexpensive.

DESCRIPTION OF DRAWING(S) - The figure shows isometric plan view of the bipolar plate.

Bipolar plate 50

Support plate 55

Intersection 56

Electrically conductive foils 58,62

Electrical contact 64

Dwg.4/5

FS EPI

FA AB; GI

MC EPI: X16-C01C; X16-C15; X16-E06A

=>

G. Gottesfeld

L92 ANSWER 1 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 2002:889103 CAPLUS
DN 137:372577
TI Methods and apparatus for a **pressure** driven methanol
fuel cell system
IN **Gottesfeld, Shimson**
PA MTI Microfuel Cells, Inc., USA
SO PCT Int. Appl., 49 pp.
CODEN: PIXXD2
DT Patent
LA English
IC ICM H01M008-04
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 47
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2002093675	A2	20021121	WO 2002-US15086	20020514
	WO 2002093675	A3	20030306		
	W: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZM, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM RW: GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR, BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG				
PRAI	US 2003031907	A1	20030213	US 2001-855982	20010515
AB	A fuel cell system including a housing defining an anode chamber and a cathode chamber and including a catalyst, a protonically conductive, but electronically nonconductive membrane positioned between the anode chamber and the cathode chamber and a first vent, a fuel chamber in gaseous communication with the anode chamber via a first valve, a liq. chamber in gaseous communication with the anode chamber via a second valve, and a mixing chamber having a second vent. The mixing chamber is in gaseous communication with the anode chamber via a third valve and receives fuel from the fuel chamber through a fuel valve, liq. from the liq. chamber via a liq. valve, and liq. effluent from the anode chamber via a liq. effluent valve. The mixing chamber also provides a fuel mixt. to the anode chamber via a fuel mixt. valve. Using effluent gases, the present invention drives fluids between elements of the fuel cell system.				
ST	methanol fuel cell system pressure driven				
IT	Catalysts (electrocatalysts; methods and app. for pressure driven methanol fuel cell system)				
IT	Fuel cells Mixing (methods and app. for pressure driven methanol fuel cell system)				
IT	67-56-1, Methanol, uses RL: PEP (Physical, engineering or chemical process); PYP (Physical process); TEM (Technical or engineered material use); PROC (Process); USES (Uses) (methods and app. for pressure driven methanol fuel cell system)				
RN	67-56-1				

L92 ANSWER 2 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 2002:80262 CAPLUS
DN 136:219496
TI System issues for Nafion-based portable direct methanol **fuel cells**
AU Pivoval, Bryan S.; Hickner, Michael; Zawodzinski, Thomas A., Jr.; Ren, Xiaoming; Gottesfeld, Shimshon; Neutzler, Jay
CS Fuel Cells and Electrochemistry, MST-11, Los Alamos National Laboratory, Los Alamos, NM, 87545, USA
SO Proceedings - Electrochemical Society (2001), 2001-4(Direct Methanol Fuel Cells), 221-230
CODEN: PESODO; ISSN: 0161-6374
PB Electrochemical Society
DT Journal
LA English
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
AB Direct methanol **fuel cells** are currently being investigated for a no. of different portable power applications, and to date Nafion has been the primary membrane used in such devices. The operating conditions for such devices (low flow rates, ambient air **pressure** and humidification, and room temp. operation) have not been extensively studied. Therefore, in an effort to better understand the behavior of Nafion-based direct methanol **fuel cells** as portable power devices, **fuel cell** expts. have been run under conditions assocd. with passive devices in conventional single cell hardware. Factors such as device size and fuel efficiency are discussed, along with water management issues for such a cell.
ST direct methanol **fuel cell** Nafion membrane
IT **Fuel cells**
 (direct methanol; system issues for Nafion-based portable direct methanol **fuel cells**)
IT Polyoxyalkylenes, uses
 RL: DEV (Device component use); USES (Uses)
 (fluorine- and sulfo-contg., ionomers; system issues for Nafion-based portable direct methanol **fuel cells**)
IT Fluoropolymers, uses
 RL: DEV (Device component use); USES (Uses)
 (polyoxyalkylene-, sulfo-contg., ionomers; system issues for Nafion-based portable direct methanol **fuel cells**)
IT Ionomers
 RL: DEV (Device component use); USES (Uses)
 (polyoxyalkylenes, fluorine- and sulfo-contg.; system issues for Nafion-based portable direct methanol **fuel cells**)
IT **Fuel cell** electrolytes
 (system issues for Nafion-based portable direct methanol **fuel cells**)
IT 66796-30-3, Nafion 117
 RL: DEV (Device component use); USES (Uses)
 (system issues for Nafion-based portable direct methanol **fuel cells**)
RE.CNT 4 THERE ARE 4 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE
(1) Meyers, J; PhD Thesis, University of California 1998
(2) Moore, R; The Electrochemical Society Proceedings Series 1998, PV98-27, P388
(3) Ren, X; J Electrochem Soc 1997, V144, PL267 CAPLUS
(4) Ren, X; J Power Sources 2000, V86, P111 CAPLUS
RN 66796-30-3

L92 ANSWER 3 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 1997:621751 CAPLUS
DN 127:265507
TI Modeling of polymer electrolyte **fuel cell** performance

AU with reformatte feed streams: effects of low levels of CO in hydrogen
Springer, T.; Zawodzinski, T.; Gottesfeld, S.
CS Los Alamos National Laboratory, Los Alamos, NM, 87545, USA
SO Proceedings - Electrochemical Society (1997), 97-13(Electrode Materials
and Processes for Energy Conversion and Storage IV), 15-24
CODEN: PESODO; ISSN: 0161-6374
PB Electrochemical Society
DT Journal
LA English
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 72
AB A simple kinetic model, based on four chem. and electrochem. surface
processes taking place at a polymer electrolyte **fuel
cell** (PEFC) Pt anode catalyst exposed to hydrogen contg. low
levels of CO, reproduces well the characteristic features of PEFC
polarization curves recorded under such conditions. These features
include a low c.d. domain of good "CO tolerance", limited essentially by
the max. rate of hydrogen dissociative chemisorption on a small fraction
of the catalyst surface area free of CO. Significantly higher voltage
losses are incurred when attempting to exceed this typical limiting c.d.,
which is detd. by the partial **pressure** of CO and the cell temp.
Significant enhancement in performance is shown to result from increases
by one-two orders of magnitude in either the inverse of the equil. const.
for CO adsorption or in the marginal rate of CO electro-oxidn. at low
anodic overpotentials. Rates of CO electro-oxidn. that would be very hard
to measure (e.g., 10 nA/cm² Pt) could have a significant effect on the
magnitude of hydrogen electrooxidn. current obtainable at low voltage
loss. X
ST modeling polymer electrolyte **fuel cell** performance;
reformatte feed stream **fuel cell**; carbon monoxide
effect **fuel cell** performance
IT **Fuel cells**
Oxidation, electrochemical
Simulation and Modeling, physicochemical
(modeling of polymer electrolyte **fuel cell**
performance with reformatte feed streams and effects of low levels of CO
in hydrogen)
IT Petroleum products
Petroleum products
Petroleum reforming
Petroleum reforming
(reformates; modeling of polymer electrolyte **fuel
cell** performance with reformatte feed streams and effects of low
levels of CO in hydrogen)
IT 7440-06-4, Platinum, uses
RL: CAT (Catalyst use); DEV (Device component use); USES (Uses)
(anode catalyst; modeling of polymer electrolyte **fuel
cell** performance with reformatte feed streams and effects of low
levels of CO in hydrogen)
IT 630-08-0, Carbon monoxide, miscellaneous
RL: MSC (Miscellaneous)
(modeling of polymer electrolyte **fuel cell**
performance with reformatte feed streams and effects of low levels of CO
in hydrogen)
IT 1333-74-0, Hydrogen, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(modeling of polymer electrolyte **fuel cell**
performance with reformatte feed streams and effects of low levels of CO
in hydrogen)
RN 7440-06-4
RN 630-08-0
RN 1333-74-0

L92 ANSWER 4 OF 7 CAPLUS COPYRIGHT 2003 ACS
AN 2000:218901 CAPLUS
DN 132:281483
TI Recent advances in direct methanol **fuel cells** at Los Alamos National Laboratory
AU Ren, X.; Zelenay, P.; Thomas, S.; Davey, J.; **Gottesfeld, S.**
CS Los Alamos National Laboratory, Los Alamos, NM, 87545, USA
SO Journal of Power Sources (2000), 86(1-2), 111-116
CODEN: JPSODZ; ISSN: 0378-7753
PB Elsevier Science S.A.
DT Journal; General Review
LA English
CC 52-0 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 51, 67
AB A review with 7 refs. This paper describes recent advances in the science and technol. of direct methanol **fuel cells** (DMFCs) made at Los Alamos National Lab. (LANL). The effort on DMFCs at LANL includes work devoted to portable power applications, funded by the Defense Advanced Research Project Agency (DARPA), and work devoted to potential transport applications, funded by the US DOE. We describe recent results with a new type of DMFC stack hardware that allows to lower the pitch per cell to 2 mm while allowing low air flow and air **pressure** drops. Such stack technol. lends itself to both portable power and potential transport applications. Power densities of 300 W/l and 1 kW/l seem achievable under conditions applicable to portable power and transport applications, resp. DMFC power system anal. based on the performance of this stack, under conditions applying to transport applications (joint effort with U.C. Davis), has shown that, in terms of overall system efficiency and system packaging requirements, a power source for a passenger vehicle based on a DMFC could compete favorably with a hydrogen-fueled **fuel cell** system, as well as with **fuel cell** systems based on fuel processing on board. As part of more fundamental studies performed, we describe optimization of anode catalyst layers in terms of PtRu catalyst nature, loading and catalyst layer compn. and structure. We specifically show that, optimized content of recast ionic conductor added to the catalyst layer is a sensitive function of the nature of the catalyst. Other elements of membrane/electrode assembly (MEA) optimization efforts are also described, highlighting our ability to resolve, to a large degree, a well-documented problem of polymer electrolyte DMFCs, namely "methanol crossover". This was achieved by appropriate cell design, enabling fuel utilization as high as 90% in highly performing DMFCs.
ST review direct methanol **fuel cell** LANL
IT Fuel cells
 (direct-methanol; recent advances in direct methanol **fuel cells** at Los Alamos National Lab.)
IT Fuel cell electrolytes
 (membrane electrolytes; recent advances in direct methanol **fuel cells** at Los Alamos National Lab.)
IT Electric vehicles
 Fuel cell anodes
 Ionic conductors
 (recent advances in direct methanol **fuel cells** at Los Alamos National Lab.)
IT 7440-06-4, Platinum, uses 7440-18-8, Ruthenium, uses
RL: CAT (Catalyst use); USES (Uses)
 (recent advances in direct methanol **fuel cells** at Los Alamos National Lab.)
IT 67-56-1, Methanol, uses
RL: NUU (Other use, unclassified); TEM (Technical or engineered material use); USES (Uses)
 (recent advances in direct methanol **fuel cells** at Los Alamos National Lab.)

RE.CNT 7 THERE ARE 7 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Chu, D; J Electrochem Soc 1994, V141, P1770 CAPLUS
- (2) Gottesfeld, S; Advances in Electrochemistry and Electrochemical Engineering 1997, V5 CAPLUS
- (3) Gottesfeld, S; Energy Storage Systems for Electronics, in press
- (4) Ren, X; J Electrochem Soc 1996, V143, PL12 CAPLUS
- (5) Ren, X; J Electrochem Soc 1997, PL267 CAPLUS
- (6) Ren, X; Proton Conducting Membrane Fuel Cells I 1995, P252 CAPLUS
- (7) Thomas, S; J Electrochem Soc, in press

RN 7440-06-4

RN 7440-18-8

RN 67-56-1

L92 ANSWER 5 OF 7 CAPLUS COPYRIGHT 2003 ACS

AN 1995:845448 CAPLUS

DN 123:261660

TI PEM **fuel cell** stack development based on membrane-electrode assemblies of ultra-low platinum loading
AU Zawodzinski, Christine; Wilson, Mahlon S.; **Gottesfeld, Shimshon**
CS Mater. Sci. Technol. Div., Los Alamos Natl. Lab., Los Alamos, NM, 87545, USA

SO Proceedings - Electrochemical Society (1995), 95-23(Proton Conducting Membrane Fuel Cells I), 57-65
CODEN: PESODO; ISSN: 0161-6374

PB Electrochemical Society

DT Journal

LA English

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)

AB Scale-up of single cell technol., based on ultra-low platinum loadings, was attempted to develop a polymer electrolyte membrane (PEM) **fuel cell** stack for stationary power generation. Initial work on scale-up to a manifolded single cell based on a 100 cm² active area is described, with the intention of combining multiples of such cells to create stacks. The cells, which are fed by **pressurized** H and air, utilize membranes catalyzed by ultra-low platinum loadings (0.14 mg Pt/cm²) and metal serpentine channel or screen flow-fields to minimize costs, while maintaining desirable power d. Some initial promising results from testing of stainless steel screens as flow-fields in such cells are given; power of 0.5 W/cm² at 0.7 V is achieved under mild flow and **pressurization** conditions.

ST polymer electrolyte membrane **fuel cell** stack

IT **Fuel cells**

(development of polymer electrolyte membrane **fuel cell** stack based on membrane-electrode assemblies of ultra-low platinum loading)

IT 7440-06-4, Platinum, uses

RL: CAT (Catalyst use); DEV (Device component use); USES (Uses) (catalytic electrodes; development of polymer electrolyte membrane **fuel cell** stack based on membrane-electrode assemblies of ultra-low platinum loading)

RN 7440-06-4

L92 ANSWER 6 OF 7 WPIDS (C) 2003 THOMSON DERWENT

AN 2001-441752 [47] WPIDS

DNN N2001-326769 DNC C2001-133497

TI **Fuel cell** for generating electrical energy, has anode distribution plate at anode for distributing methanol on its surface, and conductive sheet between plate and membrane surface forming mass transport barrier.

DC L03 X16

IN GOTTESFELD, S; REN, X

PA (REGC) UNIV CALIFORNIA

CYC 90
PI WO 2001048853 A1 20010705 (200147)* EN 21p H01M008-04
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CR CU CZ DE DK DM EE ES
FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ LC LK LR LS
LT LU LV MA MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL
TJ TM TR TT TZ UA UG UZ VN YU ZW
US 6296964 B1 20011002 (200160) H01M008-02
AU 2001024264 A 20010709 (200164) H01M008-04
ADT WO 2001048853 A1 WO 2000-US32600 20001130; US 6296964 B1 US 1999-472387
19991223; AU 2001024264 A AU 2001-24264 20001130
FDT AU 2001024264 A Based on WO 200148853
PRAI US 1999-472387 19991223
IC ICM H01M008-02; H01M008-04
ICS H01M008-10
AB WO 200148853 A UPAB: 20010822
NOVELTY - **Fuel cell** comprises anode distribution plate
at anode for distributing methanol over anodic membrane surface of
fuel cell, and conductive sheet. The conductive sheet is
provided between anode distribution plate and membrane surface forming
mass transport barrier having specified ratio of methanol current density
(J_{lim,b}) to design current (J_{cell}), of 1.1-1.3.
DETAILED DESCRIPTION - The **fuel cell** which uses
methanol direct comprises methanol feed at anode and oxygen or air feed at
cathode. The distribution plate is a corrugated plate having perforations
(12). Methanol is distributed uniformly along vertical and lateral
directions of anodic membrane surface. The **fuel cell**
further comprises cathode distribution plate at cathode for distributing
oxygen or air and for removing cathode reaction products vertically and
laterally over cathodic membrane surface. 
An INDEPENDENT CLAIM is also included for utilization of methanol in
direct methanol **fuel cell**.
USE - For generating electrical energy.
ADVANTAGE - The reactants are uniformly distributed over the active
surface of associated backing plate with small **pressure** drop
across the flow channel device. Fuel efficiency of 90%, using properly
chosen anode backing, is achieved in direct methanol **fuel cell**.
DESCRIPTION OF DRAWING(S) - The figures show fabrication of flow
field plate.
Perforation 12
1A, 1B/10
FS CPI EPI
FA AB; GI
MC CPI: L03-E04
EPI: X16-C01; X16-C09
L92 ANSWER 7 OF 7 WPIDS (C) 2003 THOMSON DERWENT
AN 2002-682608 [73] WPIDS
DNN N2002-538987 DNC C2002-192473
TI Inks for making anode and cathode catalysts for ionomer membranes of
direct methanol fuels, includes platinum and platinum-ruthenium catalysts,
purified water, and perfluorosulfonic acid ionomer.
DC A85 G02 J04 L03 X16
IN DAVEY, J; GOTTESFELD, S; REN, X; THOMAS, S C; ZELENAY, P
PA (REGC) UNIV CALIFORNIA
CYC 94
PI WO 2002045188 A2 20020606 (200273)* EN 20p H01M000-00
RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW MZ
NL OA PT SD SE SL SZ TR TZ UG ZW
W: AE AG AL AM AT AU AZ BA BB BG BR BY BZ CA CH CN CO CR CU CZ DE DK
DM DZ EE ES FI GB GD GE GH GM HR HU ID IL IN IS JP KE KG KP KR KZ

LC LK LR LS LT LU LV MA MD MG MK MN MW MX MZ NO NZ PL PT RO RU SD
SE SG SI SK SL TJ TM TR TT TZ UA UG UZ VN YU ZA ZW

AU 2002039439 A 20020611 (200273) H01M000-00

ADT WO 2002045188 A2 WO 2001-US45619 20011031; AU 2002039439 A AU 2002-39439
20011031

FDT AU 2002039439 A Based on WO 200245188

PRAI US 2000-715211 20001114

IC ICM H01M000-00

AB WO 200245188 A UPAB: 20021113

NOVELTY - Inks comprise a platinum catalyst for a cathode catalyst and a platinum-ruthenium catalyst for an anode catalyst, purified water which is 4-20 times that of the catalyst by weight, and a perfluorosulfonic acid ionomer in an amount effective to provide an ionomer content in dried anode and cathode catalyst coats of 20-80 vol. %.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for the following:

(a) the production of inks for use in anode and cathode catalysts applied to membranes for direct methanol **fuel cells**, comprising combining platinum or platinum-ruthenium catalysts with purified water to form a first mixture, cooling the first mixture to a temperature that reduces evaporation of water from the first mixture, placing the first mixture in a cooler and agitating the mixture to obtain a first homogeneous suspension, adding an alcoholic solution of perfluorosulfonic acid to the first homogeneous suspension to provide a second mixture, agitating the second mixture in the cooler to obtain a homogeneous ink suitable for application to the appropriate anode or cathode side of the membrane; and

(b) a method for applying anode and cathode catalyst to anode and cathode surfaces for a direct methanol **fuel cell**, comprising forming a mixture of water, perfluorosulfonic acid ionomer, alcohols, and catalyst of platinum-ruthenium for the anode or platinum catalyst for the cathode while cooling and agitating the solution, placing the mixture in a cooler, continuously agitating the mixture, spraying the mixture on the anode or cathode surfaces as determined by the catalyst content.

USE - For making anode and cathode catalysts for application to ionomer membranes for direct methanol fuels.

Dwg. 0/5

FS CPI EPI

FA AB

MC CPI: A12-E06; G02-A04A; J04-E04; L03-E04

EPI: X16-C; X16-C09

=>

Wilson

L105 ANSWER 1 OF 4 CAPLUS COPYRIGHT 2003 ACS
AN 2000:145143 CAPLUS
DN 132:154467
TI Ambient pressure fuel cell system
IN Wilson, Mahlon S.
PA The Regents of the University of California, USA
SO PCT Int. Appl., 28 pp.
CODEN: PIXXD2
DT Patent
LA English
IC ICM H01M008-04
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 2000011745	A1	20000302	WO 1999-US17573	19990803
	W: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, GM, HR, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM RW: GH, GM, KE, LS, MW, SD, SL, SZ, UG, ZW, AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG				
	US 6117577	A	20000912	US 1998-135965	19980818
	CA 2340765	AA	20000302	CA 1999-2340765	19990803
	AU 9957717	A1	20000314	AU 1999-57717	19990803
	EP 1110264	A1	20010627	EP 1999-945013	19990803
	EP 1110264	B1	20030402		
	R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, SI, LT, LV, FI, RO				
	AT 236457	E	20030415	AT 1999-945013	19990803
PRAI	US 1998-135965	A	19980818		
	WO 1999-US17573	W	19990803		
AB	An ambient pressure fuel cell system is provided with a fuel cell stack formed from a plurality of fuel cells having membrane/electrode assemblies (MEAs) that are hydrated with liq. water and bipolar plates with anode and cathode sides for distributing hydrogen fuel gas and water to a first side of each one of the MEAs and air with reactant oxygen gas to a second side of each one of the MEAs. A pump supplies liq. water to the fuel cells. A recirculating system may be used to return unused hydrogen fuel gas to the stack. A near-ambient pressure blower blows air through the fuel cell stack in excess of reaction stoichiometric amts. to react with the hydrogen fuel gas. X				
ST	fuel cell system ambient pressure				
IT	Fuel cells (ambient pressure fuel cell system)				
IT	Epoxy resins, uses RL: DEV (Device component use); USES (Uses) (ambient pressure fuel cell system)				
IT	7782-42-5, Graphite, uses RL: DEV (Device component use); USES (Uses) (ambient pressure fuel cell system)				
IT	7782-44-7, Oxygen, reactions RL: RCT (Reactant); RACT (Reactant or reagent) (ambient pressure fuel cell system)				
IT	1333-74-0, Hydrogen, uses RL: TEM (Technical or engineered material use); USES (Uses) (ambient pressure fuel cell system)				

RE.CNT 8 THERE ARE 8 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE

- (1) Cargnelli; US 5753383 A 1998 CAPLUS
- (2) Ernst; US 5912088 1999
- (3) Farooque; US 5084362 A 1992 CAPLUS
- (4) Nakazawa; US 5134043 A 1992 CAPLUS
- (5) Okamoto; US 5837393 A 1998
- (6) Reiser; US 5853909 A 1998 CAPLUS
- (7) Scheffler; US 4859545 A 1989
- (8) Sederquist; US 5330857 A 1994 CAPLUS

RN 7782-42-5
RN 7782-44-7
RN 1333-74-0

L105 ANSWER 2 OF 4 CAPLUS COPYRIGHT 2003 ACS
AN 1995:845448 CAPLUS
DN 123:261660
TI PEM fuel cell stack development based on membrane-electrode assemblies of ultra-low platinum loading
AU Zawodzinski, Christine; Wilson, Mahlon S.; Gottesfeld, Shimshon
CS Mater. Sci. Technol. Div., Los Alamos Natl. Lab., Los Alamos, NM, 87545, USA
SO Proceedings - Electrochemical Society (1995), 95-23(Proton Conducting Membrane Fuel Cells I), 57-65
CODEN: PESODO; ISSN: 0161-6374
PB Electrochemical Society
DT Journal
LA English
CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
AB Scale-up of single cell technol., based on ultra-low platinum loadings, was attempted to develop a polymer electrolyte membrane (PEM) fuel cell stack for stationary power generation. Initial work on scale-up to a manifolded single cell based on a 100 cm² active area is described, with the intention of combining multiples of such cells to create stacks. The cells, which are fed by pressurized H and air, utilize membranes catalyzed by ultra-low platinum loadings (0.14 mg Pt/cm²) and metal serpentine channel or screen flow-fields to minimize costs, while maintaining desirable power d. Some initial promising results from testing of stainless steel screens as flow-fields in such cells are given; power of 0.5 W/cm² at 0.7 V is achieved under mild flow and pressurization conditions.
ST polymer electrolyte membrane fuel cell stack
IT Fuel cells
(development of polymer electrolyte membrane fuel cell stack based on membrane-electrode assemblies of ultra-low platinum loading)

IT 7440-06-4, Platinum, uses
RL: CAT (Catalyst use); DEV (Device component use); USES (Uses)
(catalytic electrodes; development of polymer electrolyte membrane fuel cell stack based on membrane-electrode assemblies of ultra-low platinum loading)
RN 7440-06-4

L105 ANSWER 3 OF 4 CAPLUS COPYRIGHT 2003 ACS
AN 1993:84416 CAPLUS
DN 118:84416
TI Membrane catalyst layer for fuel cells
IN Wilson, Mahlon Scott
PA United States Dept. of Energy, USA
SO PCT Int. Appl., 29 pp.
CODEN: PIXXD2
DT Patent
LA English

IC ICM H01M008-10

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 38

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	WO 9215121	A1	19920903	WO 1992-US1058	19920218
	W: CA, JP				
	RW: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LU, MC, NL, SE				
	US 656329	A0	19930501	US 1991-656329	19910219
	US 5234777	A	19930810	US 1991-736876	19910729
	JP 05507583	T2	19931028	JP 1992-507039	19920218
	EP 600888	A1	19940615	EP 1992-907218	19920218
	EP 600888	B1	19970827		
	R: DE, FR, GB				

PRAI US 1991-656329 19910219
US 1991-736876 19910729
WO 1992-US1058 19920218

AB *Fuel cells* incorporate a .ltorsim.10-.mu.m catalyst layer between a solid polymer electrolyte membrane and a porous electrode backing. The catalyst layer has C-supported Pt catalyst loading >0.1 and .ltorsim.0.35 mg Pt/cm². The layer is formed as an ink that is spread and cured on a film-release blank. The cured film is transferred to the membrane and hot **pressed** into the surface to form a catalyst layer having a controlled thickness and catalyst distribution. Alternately, the catalyst layer is formed by applying a Na⁺ form of a perfluorosulfonate ionomer directly to the membrane, drying the film at .gtoreq.150.degree., and then converting the film back to the protonated form of the ionomer. The layer has an adequate gas permeability so that the cell performance is not affected and has an effective d. and particle distribution to optimize proton access to the catalyst and electronic continuity for electron flow from the half-cell reaction occurring at the catalyst.

ST ionomer perfluorosulfonate platinum electrode; **fuel cell** platinum electrode

IT Ionomers

RL: USES (Uses)

(fluoropolymers, sulfo-contg., protonated, electrodes from platinum-contg., catalytic, for **fuel cells**)

IT Electrodes

(fuel-cell, catalytic, platinum in protonated perfluorosulfonate ionomer)

IT Fluoropolymers

RL: USES (Uses)

(ionomers, sulfo-contg., protonated, electrodes from platinum-contg., catalytic, for **fuel cells**)

IT 7440-06-4, Platinum, uses

RL: USES (Uses)

(electrodes, catalytic, in protonated perfluorosulfonate ionomer, for **fuel cells**)

RN 7440-06-4

L105 ANSWER 4 OF 4 CAPLUS COPYRIGHT 2003 ACS

AN 2001:480625 CAPLUS

DN 135:63840

TI Pleated metal bipolar assembly for use in electrochemical cells

IN Wilson, Mahlon S.; Zawodzinski, Christine

PA The Regents of the University of California, USA

SO U.S., 8 pp.

CODEN: USXXAM

DT Patent

LA English

IC ICM H01M002-14

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NCL 429038000

CC 52-2 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 72

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 6255012	B1	20010703	US 1999-444216	19991119
PRAI	US 1999-444216		19991119		
AB	A thin low-cost bipolar plate for an electrochem. cell is formed from a polymer support plate with first flow channels on a first side of the support plate and second flow channels on a second side of the support plate, where the first flow channels and second flow channels have intersecting locations and have a depth effective to form openings through the support plate at the intersecting locations. A first foil of elec. conductive material is pressed into the first flow channels. A second foil of elec. conductive material pressed into the second flow channels so that elec. contact is made between the first and second foils at the openings through the support plate. A particular application of the bipolar plate is in polymer electrolyte fuel cells.				
ST	fuel cell pleated metal bipolar assembly; electrochem cell pleated metal bipolar assembly				
IT	Fuel cells (pleated metal bipolar assembly for use in electrochem. cells)				
IT	Polymers, uses RL: TEM (Technical or engineered material use); USES (Uses) (support plate; pleated metal bipolar assembly for use in electrochem. cells)				

RE.CNT 6 THERE ARE 6 CITED REFERENCES AVAILABLE FOR THIS RECORD
RE

- (1) Malhi; US 5789093 1998 CAPLUS
- (2) Mukohyama; US 5798188 1998 CAPLUS
- (3) Plowman; US 4755272 1988 CAPLUS
- (4) Shinn; US 3880670 1975 CAPLUS
- (5) Tajima; US 5541015 1996 CAPLUS
- (6) Wilson; US 5798187 1998 CAPLUS

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429|NC

L111 ANSWER 1 OF 2 USPATFULL

AN 95:80009 USPATFULL
TI Method for making reinforced ion exchange membranes
IN Banerjee, Shoibal, Newark, DE, United States
PA E. I. Du Pont de Nemours and Company, Wilmington, DE, United States
(U.S. corporation)
PI US 5447636 19950905
AI US 1993-168869 19931214 (8)
DT Utility
FS Granted
EXNAM Primary Examiner: Kim, John
CLMN Number of Claims: 19
ECL Exemplary Claim: 1
DRWN No Drawings
LN.CNT 684

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB A method for making a composite fluorinated ion exchange membrane is provided comprising applying a continuous fluorinated ion exchange polymer film to a porous reinforcing substrate, preferably made from a polyolefin such as linear high density polyethylene. The composite membrane is made by applying a solvent to the surface of the film and/or the substrate and then affixing the reinforcing substrate. These reinforced membranes may advantageously be used as separators in fuel cells, and as selective barriers in permeation separation and facilitated transport operations.

L111 ANSWER 2 OF 2 USPATFULL

AN 88:14606 USPATFULL
TI Reinforced flexible graphite sheet
IN Fukuda, Hiroyuki, Iwaki, Japan
Shigeta, Masatomo, Iwaki, Japan
Kaji, Hisatsugu, Iwaki, Japan
Saitoh, Kuniyuki, Abiko, Japan
PA Kureha Kagaku Kogyo Kabushiki Kaisha, Tokyo, Japan (non-U.S. corporation)
PI US 4729910 19880308
AI US 1985-719562 19850403 (6)
PRAI JP 1984-71659 19840410
DT Utility
FS Granted
EXNAM Primary Examiner: Lusignan, Michael R.
LREP Cushman, Darby & Cushman
CLMN Number of Claims: 7
ECL Exemplary Claim: 1
DRWN 1 Drawing Figure(s); 1 Drawing Page(s)
LN.CNT 219

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB Disclosed herein is a reinforced flexible graphite sheet uniformly containing a thermosetting resin prepared by impregnating a flexible graphite sheet with a liquid thermosetting resin having the carbonizing yield of more than 20% under a reduced pressure and thereafter, heat-setting the resin, an amount of the resin impregnated being in the range of 0.5 to 20% by weight based on the flexible graphite sheet.

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429/NCL

L124 ANSWER 1 OF 13 USPATFULL
AN 2001:208280 USPATFULL
TI Membrane electrode assembly
IN Debe, Mark K., Stillwater, MN, United States
Poirier, Richard J., White Bear Lake, MN, United States
Wackerfuss, Michael K., Roseville, MN, United States
Ziegler, Raymond J., Glenwood City, WI, United States
PA 3M Innovative Properties Company, Saint Paul, MN, United States (U.S.
corporation)
PI US 6319293 B1 20011120
AI US 1998-208657 19981210 (9)
RLI Division of Ser. No. US 1997-948599, filed on 10 Oct 1997, now patented,
Pat. No. US 5879828
DT Utility
FS GRANTED
EXNAM Primary Examiner: Bell, Bruce F.
LREP Dahl, Philip Y.
CLMN Number of Claims: 7
ECL Exemplary Claim: 1
DRWN 25 Drawing Figure(s); 23 Drawing Page(s)
LN.CNT 1921
CAS INDEXING IS AVAILABLE FOR THIS PATENT.
AB A membrane electrode assembly is provided comprising an ion conducting
membrane and one or more electrode layers that comprise nanostructured
elements, wherein the nanostructured elements are in incomplete contact
with the ion conducting membrane. This invention also provides methods
to make the membrane electrode assembly of the invention. The membrane
electrode assembly of this invention is suitable for use in
electrochemical devices, including proton exchange membrane **fuel**
cells, electrolyzers, chlor-alkali separation membranes, and the
like.

X

L124 ANSWER 2 OF 13 USPATFULL
AN 1999:30515 USPATFULL
TI Membrane electrode assembly
IN Debe, Mark K., Stillwater, MN, United States
Poirier, Richard J., White Bear Lake, MN, United States
Wackerfuss, Michael K., Roseville, MN, United States
Ziegler, Raymond J., Glenwood City, WI, United States
PA Minnesota Mining and Manufacturing Company, St. Paul, MN, United States
(U.S. corporation)
PI US 5879828 19990309
AI US 1997-948599 19971010 (8)
DT Utility
FS Granted

EXNAM Primary Examiner: Bell, Bruce F.
LREP Dahl, Philip Y.
CLMN Number of Claims: 14
ECL Exemplary Claim: 1
DRWN 25 Drawing Figure(s); 23 Drawing Page(s)
LN.CNT 1974
CAS INDEXING IS AVAILABLE FOR THIS PATENT.

✓

AB A membrane electrode assembly is provided comprising an ion conducting
membrane and one or more electrode layers that comprise nanostructured
elements, wherein the nanostructured elements are in incomplete contact
with the ion conducting membrane. This invention also provides methods
to make the membrane electrode assembly of the invention. The membrane
electrode assembly of this invention is suitable for use in
electrochemical devices, including proton exchange membrane **fuel**
cells, electrolyzers, chlor-alkali separation membranes, and the
like.

L124 ANSWER 3 OF 13 USPATFULL

AN 2002:3770 USPATFULL
TI Membrane electrode assembly, and solid polymer fuel cell using the assembly
IN Tsusaka, Kyoko, Aichi-ken, JAPAN
Kawasumi, Masaya, Aichi-ken, JAPAN
Morimoto, Yu, Aichi-ken, JAPAN
PA Kabushiki Kaisha Toyota Chuo Kenkyusho, Aichi-gun, JAPAN (non-U.S. corporation)
PI US 2002001744 A1 20020103
AI US 2001-858979 A1 20010517 (9)
PRAI JP 2000-145877 20000518
JP 2000-401275 20001228
DT Utility
FS APPLICATION
LREP OBLON SPIVAK MCCLELLAND MAIER & NEUSTADT PC, FOURTH FLOOR, 1755 JEFFERSON DAVIS HIGHWAY, ARLINGTON, VA, 22202
CLMN Number of Claims: 18
ECL Exemplary Claim: 1
DRWN 6 Drawing Page(s)
LN.CNT 1844
CAS INDEXING IS AVAILABLE FOR THIS PATENT.
AB Disclosed is a membrane electrode assembly obtained by bonding electrodes to both surfaces of a solid polymer electrolyte membrane suitably for use in a solid polymer fuel cell. In order to maintain not only the solid polymer electrolyte but also the electrode in appropriate wet states, the catalyst layer of the assembly contains a metallocxane polymer in the intra-catalyst-layer electrolyte including an electrode catalyst preferably in an amount of 0.5 to 50 wt % of the total weight of the intra-catalyst-layer electrolyte and the metallocxane polymer contained therein exclusive of the electrode catalyst. It is also preferred that a metallocxane polymer be included in the solid polymer electrolyte membrane in an amount of 0.5 to 50 wt % of the total weight of the solid polymer electrolyte membrane and the metallocxane polymer contained therein. The membrane electrode assembly constituted as above may be obtained by bonding the catalyst layer containing metallocxane monomer to the solid polymer electrolyte membrane by hot-pressing or the like. Here, it is preferred that the gas-phase surface of the intra-catalyst-layer electrolyte be covered with a gas-permeable, water repellent layer. The water-repellent layer may be formed by various methods such as applying hydrophobic metallocxane precursor, followed by polycondensation.

L124 ANSWER 4 OF 13 USPATFULL

AN 2002:32486 USPATFULL
TI Composite catalyst for solid polymer electrolyte type fuel cell and processes for producing the same
IN Hitomi, Shuji, Kyoto, JAPAN
Tsumura, Naohiro, Kyoto, JAPAN
Mizutani, Shunsuke, Kyoto, JAPAN
PI US 2002019308 A1 20020214
US 6492295 B2 20021210
AI US 2001-808032 A1 20010315 (9)
PRAI JP 2000-72347 20000315
JP 2000-126636 20000426
JP 2000-140483 20000512
DT Utility
FS APPLICATION
LREP SUGHRUE, MION, ZINN, MACPEAK & SEAL, PLLC, 2100 PENNSYLVANIA AVENUE, N.W., WASHINGTON, DC, 20037-3213
CLMN Number of Claims: 21
ECL Exemplary Claim: 1

DRWN 16 Drawing Page(s)

LN.CNT 981

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB A composite catalyst which comprises a catalyst particle and at least one member selected from the group consisting of a porous or net-form cation-exchange resin and a porous or net-form hydrophobic polymer, wherein the resin and polymer exist on the surface of the catalyst particle.

L124 ANSWER 5 OF 13 USPAT2

AN 2002:32486 USPAT2

TI Composite catalyst for solid polymer electrolyte type fuel cell and processes for producing the same

IN Hitomi, Shuji, Kyoto, JAPAN

Tsumura, Naohiro, Kyoto, JAPAN

Mizutani, Shunsuke, Kyoto, JAPAN

PA Japan Storage Battery Co., Ltd., Kyoto, JAPAN (non-U.S. corporation)

PI US 6492295 B2 20021210

AI US 2001-808032 20010315 (9)

PRAI JP 2000-72347 20000315

JP 2000-126636 20000426

JP 2000-140483 20000512

DT Utility

FS GRANTED

EXNAM Primary Examiner: Bell, Mark L.; Assistant Examiner: Hailey, Patricia L.

LREP Sughrue Mion, PLLC

CLMN Number of Claims: 21

ECL Exemplary Claim: 1

DRWN 25 Drawing Figure(s); 16 Drawing Page(s)

LN.CNT 979

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB A composite catalyst which comprises a catalyst particle and at least one member selected from the group consisting of a porous or net-form cation-exchange resin and a porous or net-form hydrophobia polymer, wherein the resin and polymer exist on the surface of the catalyst particle.

L124 ANSWER 6 OF 13 USPATFULL

AN 2003:10506 USPATFULL

TI Solid polymer type fuel battery

IN Ito, Hideki, Hiroshima, Hiroshima-ken, JAPAN

Kobayashi, Toshiro, Hiroshima, Hiroshima-ken, JAPAN

Moriga, Takuya, Hiroshima, Hiroshima-ken, JAPAN

Yamada, Akihiko, Yokohama Kanagawa-ken, JAPAN

PI US 2003008200 A1 20030109

AI US 2002-129133 A1 20020516 (10)

WO 2001-JP8103 20010918

PRAI JP 2000-282397 20000918

DT Utility

FS APPLICATION

LREP OBLON SPIVAK MCCLELLAND MAIER & NEUSTADT PC, FOURTH FLOOR, 1755 JEFFERSON DAVIS HIGHWAY, ARLINGTON, VA, 22202

CLMN Number of Claims: 9

ECL Exemplary Claim: 1

DRWN 6 Drawing Page(s)

LN.CNT 509

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB A polymer electrolyte fuel cell comprising a cell having a solid polymer film separators disposed on both sides of the cell so as to interpose the cell therebetween, and diffusion layers disposed between the cell and the separators and each having a substrate comprising an electrically conductive porous material and a slurry layer disposed on the substrate, wherein at least a part of each diffusion

layer is provided with a gas barrier for preventing the permeation of gas in a direction parallel to the major surface of the diffusion layer.

L124 ANSWER 7 OF 13 USPATFULL

AN 75:41112 USPATFULL
TI Gas electrodes and a process for producing them
IN Kordesch, Karl V., Lakewood, OH, United States
PA Union Carbide Corporation, New York, NY, United States (U.S. corporation)
PI US 3899354 19750812
AI US 1973-395552 19730910 (5)
DT Utility
FS Granted
EXNAM Primary Examiner: Mack, John H.; Assistant Examiner: Feeley, H. A.
LREP O'Brien, C. F.
CLMN Number of Claims: 12
ECL Exemplary Claim: 1
DRWN No Drawings
LN.CNT 586
CAS INDEXING IS AVAILABLE FOR THIS PATENT.
AB A thin catalyzed gas electrode for fuel cells comprising a porous wet proofed conductive substrate having a first water-repellent porous active conductive layer over which is a surface-deposited noble metal catalyst in an amount of at least about 0.5 mg/cm.sup.2, and a process for producing such an electrode.

L124 ANSWER 8 OF 13 USPATFULL

AN 1998:150537 USPATFULL
TI Process for forming a catalyst layer on an electrode by spray-drying
IN Tada, Tomoyuki, Kanagawa, Japan
PA Tanaka Kikinzoku Kogyo K.K., Japan (non-U.S. corporation)
Watanabe, Masahiro, Japan (non-U.S. individual)
Stonehart Associates Inc., Madison, CT, United States (U.S. corporation)
PI US 5843519 19981201
AI US 1995-543632 19951016 (8)
PRAI JP 1994-277108 19941017
JP 1994-332291 19941017
JP 1994-289288 19941028
JP 1994-289289 19941028
DT Utility
FS Granted
EXNAM Primary Examiner: Bareford, Katherine A.
LREP Klauber & Jackson
CLMN Number of Claims: 2
ECL Exemplary Claim: 1
DRWN 11 Drawing Figure(s); 5 Drawing Page(s)
LN.CNT 896
CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB Disclosed herein are four aspects of processes for preparing an electrode for a solid polymer electrolyte electrochemical cell such as a fuel cell. According to the processes, the electrode having a thin and uniform electrocatalyst layer can be obtained effectively and economically. The processes include a spraying method, a paste rolling method and a dry mixture method.

L124 ANSWER 9 OF 13 USPATFULL

AN 2001:190851 USPATFULL
TI Membrane-electrode unit for polymer electrolyte fuel cells and processes for their preparation
IN Zuber, Ralf, Grossostheim, Germany, Federal Republic of
Fehl, Knut, Schluchtern, Germany, Federal Republic of
Starz, Karl-Anton, Rodenbach, Germany, Federal Republic of
Stenke, Udo, Mainaschaff, Germany, Federal Republic of

PA Degussa AG, Hanau, Germany, Federal Republic of (non-U.S. corporation)
PI US 6309772 B1 20011030
AI US 1999-274018 19990322 (9)
PRAI DE 1998-19812592 19980323
DT Utility
FS GRANTED
EXNAM Primary Examiner: Nguyen, Nam; Assistant Examiner: VerSteeg, Steven H.
LREP Smith, Gambrell & Russell, LLP
CLMN Number of Claims: 29
ECL Exemplary Claim: 1
DRWN 5 Drawing Figure(s); 3 Drawing Page(s)
LN.CNT 578

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB A membrane-electrode unit for polymer-electrolyte **fuel cells**. The membrane-electrode unit consists of a polymer electrolyte membrane and porous reaction layers applied to both sides comprising a catalyst and a proton-conducting polymer, a so-called ionomer. The membrane-electrode unit is characterized in that one part A1 of the catalyst of the reaction layers is saturated with the ionomer and is embedded in the ionomer whereas one part A2 of the catalyst is kept free from the ionomer, where the parts A1 and A2 are in a weight ratio of 1:1 to 20:1. X

L124 ANSWER 10 OF 13 USPATFULL

AN 2000:164209 USPATFULL
TI Catalyst layer for polymer electrolyte **fuel cells**
IN Zuber, Ralf, Grossostheim, Germany, Federal Republic of
Karch, Ralf, Kleinostheim, Germany, Federal Republic of
Fehl, Knut, Schluchtern-Ramholz, Germany, Federal Republic of
Starz, Karl-Anton, Rodenbach, Germany, Federal Republic of
PA Degussa-Huls Aktiengellschaft, Frankfurt am Main, Germany, Federal
Republic of (non-U.S. corporation)
PI US 6156449 20001205
AI US 1999-376438 19990818 (9)
PRAI DE 1998-19837669 19980820
DT Utility
FS Granted
EXNAM Primary Examiner: Bell, Bruce F.
CLMN Number of Claims: 26
ECL Exemplary Claim: 1
DRWN 5 Drawing Figure(s); 2 Drawing Page(s)
LN.CNT 634

CAS INDEXING IS AVAILABLE FOR THIS PATENT. ✓

AB A catalyst layer on a substrate material which contains a proton-conducting polymer (ionomer), electrically conductive carbon particles and fine particles of at least one precious metal. The catalyst layer is obtainable by coating the substrate material with an ink which contains a dispersion of the carbon particles and at least one organic precious metal complex compound in a solution of the ionomer, and drying the coating below a temperature at which the ionomer or the substrate material is thermally damaged, the precious metals in the complex compounds being present with an oxidation number of 0 and the complex compounds being thermally decomposed during drying to form the fine precious metal particles.

L124 ANSWER 11 OF 13 USPATFULL

AN 90:44235 USPATFULL
TI Gas permeable electrode
IN Watanabe, Masahiro, No. 2-10, Kitashin 1-chome, Kofu-shi, Yamanashi,
Japan
Motoo, Satoshi, No. 5-24, Takeda 3-chome, Kofu-shi, Yamanashi, Japan
Furuya, Nagakazu, No. 4-3-31, Ohte 2-chome, Kofu-shi, Yamanashi, Japan
PA Watanabe, Masahiro, Japan (non-U.S. individual)

Motoo, Satoshi, Japan (non-U.S. individual)
Furuya, Nagakazu, Japan (non-U.S. individual)
Tanaka Kikinzoku Kogyo, Japan (non-U.S. corporation)
PI US 4931168 19900605
AI US 1989-356612 19890523 (7)
RLI Continuation of Ser. No. US 1987-22437, filed on 6 Mar 1987, now abandoned
DT Utility
FS Granted
EXNAM Primary Examiner: Niebling, John F.; Assistant Examiner: Gorgos, Kathryn
LREP Klauber & Jackson
CLMN Number of Claims: 11
ECL Exemplary Claim: 2
DRWN 52 Drawing Figure(s); 11 Drawing Page(s)
LN.CNT 995

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB Disclosed herein is a gas permeable electrode which comprises a gas permeable layer and a reaction layer, the reaction layer comprising hydrophobic portions and hydrophilic portions. Electrolyte penetrates into the reaction layer and does not penetrate into the gas permeable layer, and only the gas produced on the electrode and the gas supplied penetrate into the gas permeable layer. The gas can be released from the rear side of the electrode and the reaction surface thereof is never covered with the gas.

L124 ANSWER 12 OF 13 USPATFULL

AN 80:59004 USPATFULL
TI Method of making improved hydrogenation catalyst
IN Berchielli, Aldo S., Westerly, RI, United States
PA Chireau, Roland F., Quaker Hills, CT, United States
Yardney Electric Corporation, Pawcatuck, CT, United States (U.S. corporation)
PI US 4235748 19801125
AI US 1979-16219 19790228 (6)
DT Utility
FS Granted
EXNAM Primary Examiner: Konopka, P. E.
LREP Nist, Donald E.
CLMN Number of Claims: 9
ECL Exemplary Claim: 1,4,6
DRWN No Drawings
LN.CNT 406

CAS INDEXING IS AVAILABLE FOR THIS PATENT.

AB The present invention provides a method of making an improved catalyst useful for hydrogenation of organic compounds and in other reactions. The method comprises mixing a porous particulate refractory support material with a hydrophobic polymeric binder to form an essentially homogeneous mixture and then catalyzing this mixture by introducing into the pores of the mixture a salt of a metal of the eighth group of the Periodic Table in a non-polar solvent in which the salt is ionizable. Further in accordance with the method, the salt is reduced to the catalytic metal within the pores of the mixture, specifically within the pores of the support material, by contacting with a selected reducing agent, for example, hydrazine or sodium bis (2-methoxyethoxy) aluminum hydride, preferably in solution. A porous structure is formed from the mixture either before or after the catalysis. Preferably, the structure is sintered as a final step. The support material in the mixture is preferably activated carbon having an average particle diameter of about 0.01-0.3 microns and the structure preferably has pores of an average diameter of about 30-300 Å. Most preferably, the binder is polytetrafluoroethylene. During the mixing step, the mixture is sheared to a fine filament-like fibrous structure of a uniform spongy nature with no grossly detectable free

carbon or polytetrafluoroethylene particles.

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AN 76:49202 USPATFULL
TI Method for catalyzing a fuel cell electrode and an electrode so produced
IN Katz, Murray, Newington, CT, United States
Kaufman, Arthur, Bloomfield, CT, United States
PA United Technologies Corporation, East Hartford, CT, United States (U.S. corporation)
PI US 3979227 19760907
AI US 1976-645962 19760102 (5)
RLI Division of Ser. No. US 1974-533918, filed on 18 Dec 1974, now patented, Pat. No. US 3932197
DT Utility
FS Granted
EXNAM Primary Examiner: Mack, John H.; Assistant Examiner: Feeley, H. A.
LREP Stone, Steven F.
CLMN Number of Claims: 4
ECL Exemplary Claim: 1
DRWN 2 Drawing Figure(s); 2 Drawing Page(s)
LN.CNT 586
CAS INDEXING IS AVAILABLE FOR THIS PATENT.
AB A porous conducting particle, hydrophobic bonded, substrate supported electrode is prewetted with the electrolyte. A D.C. voltage is applied to the electrode to assist in the prewetting with the electrolyte. A soluble catalyst-containing material is then introduced into the electrode structure and the catalyst deposited within the electrode. By appropriate selection of the porous conducting particles and the catalyst-applying techniques, precise control of the location of the catalyst can be obtained. If graphite materials are used as the conducting particles, a catalyst-containing salt is allowed to dissolve in the electrolyte in the prewetted electrode, and the catalyst-containing material is reduced to the metal. If the reduction is done by reaction with a reducing gas such as hydrogen, the catalyst will be deposited only in those regions of the electrode at which there is an electrolyte-reactant gas interface which is in electrical-conducting relationship with the substrate. Alternatively, extremely precise amounts of catalyst can be deposited within the electrode structure by use of a solution of a compound of the catalyst whose wettability with the hydrophobic material varies as the solution evaporates. By this technique almost 100% of the catalyst can be deposited within the electrode structure on the hydrophilic region, with virtually no losses in the hydrophobic material.

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